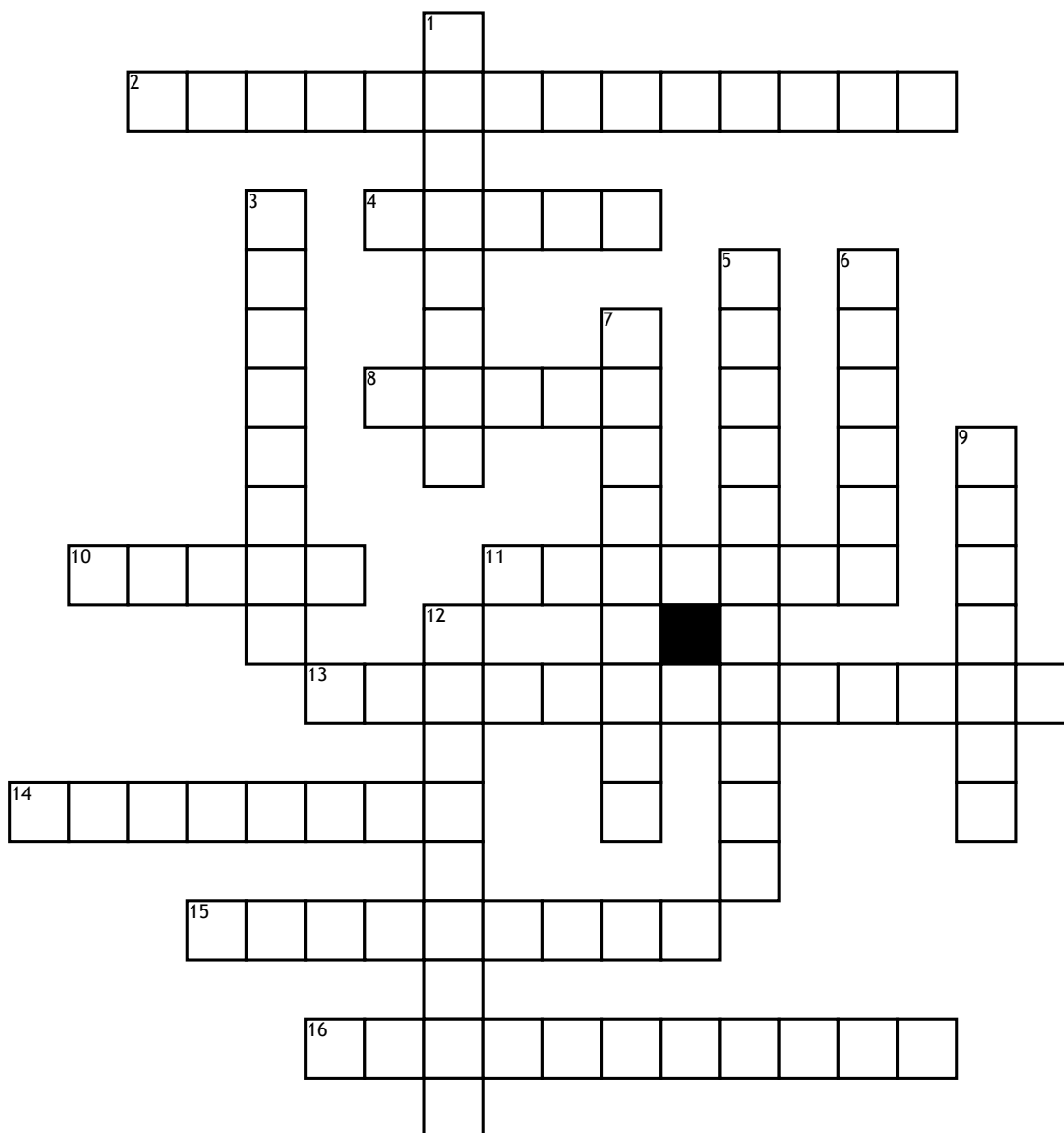


Name: _____

Date: _____

AP Chemistry Escape Room Review



Across

2. London dispersion forces increases as _____ increases

_____ increases

4. Determines if a reaction is thermodynamically feasible

8. Location for oxidation in an electrochemical cell

10. The order of a reactant that doubles the rate, when its concentration is doubled

11. Atom with greatest atomic radius in period 2

13. Beers law states that the _____ of a solution is proportional to its absorbance

14. A molecule with polar bonds and an symmetric shape

15. An atom with the electron configuration of $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$

16. This will form when $Q > K_{sp}$

Down

1. The _____ reactant is used up first and determine the theoretical yield

3. When Q decreases E_{cell} will _____

5. If asked to determine which direction a reaction will proceed to reach equilibrium I should compare the reaction quotient to an _____

constant.

6. Which noble gas would effuse the fastest at 25C?

7. When K_{eq} is less than 1, would you expect reactants or products to be present at a greater concentration at equilibrium?

9. The driving force for a reaction that is only thermodynamically feasible at high temperature

12. Given the reaction, $N_2O_4(g) \leftrightarrow 2NO_2(g)$, entropy _____