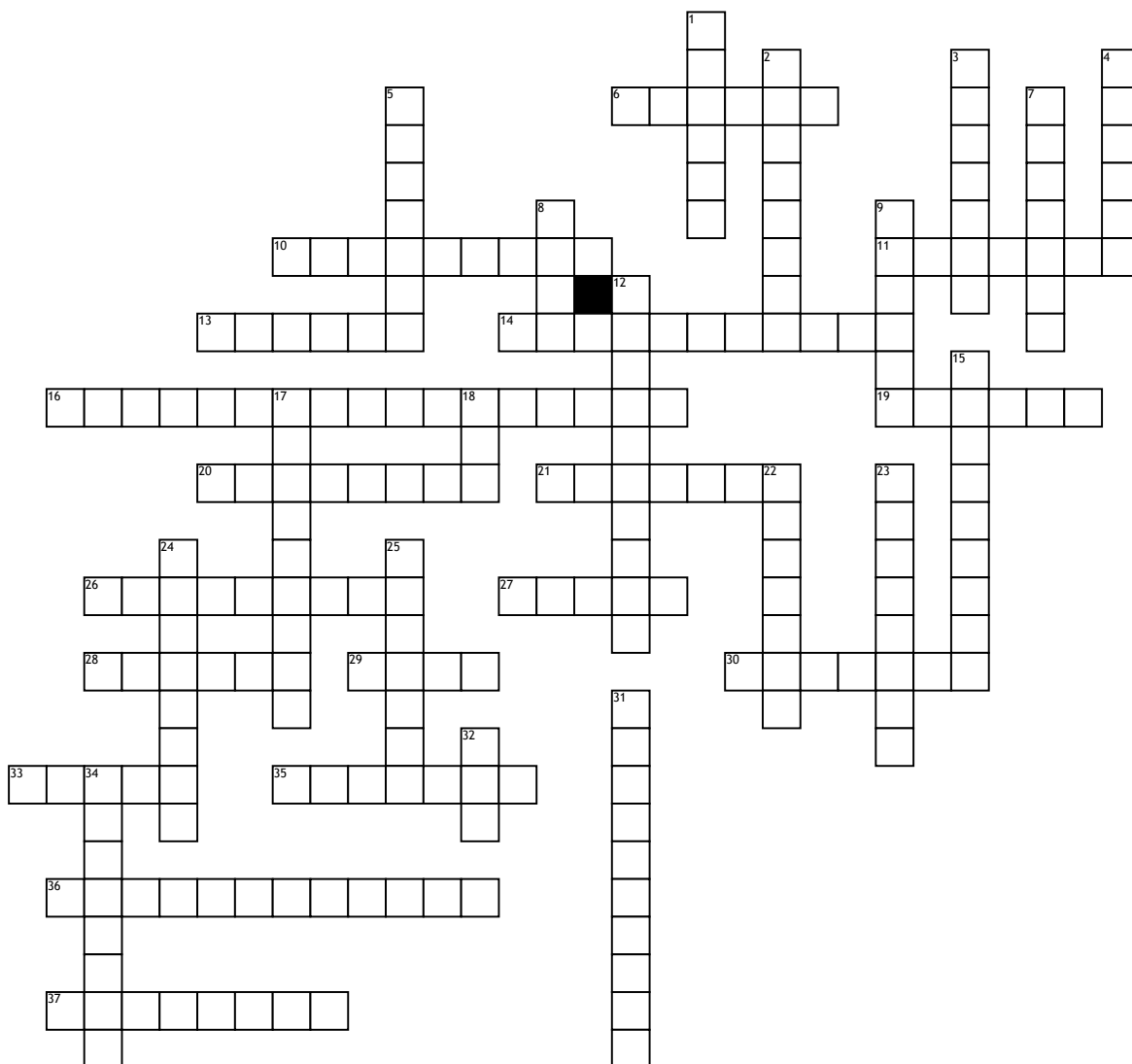


Name: \_\_\_\_\_

Date: \_\_\_\_\_

# ATOMIC STRUCTURE



## Across

6. has a mass of about 1 amu  
 10. what is mass number of Cl-36  
 11. When an atom has equal protons and electrons it is \_\_\_\_\_.  
 13. An atom has 20 protons and 18 electrons it is a \_\_\_\_\_.  
 14. atom that gains 2 electrons would have this charge  
 16. average mass of all isotopes  
 19. atomic number of 47  
 20. negative subatomic particle  
 21. subatomic particle with no charge  
 26. In a neutral atom of Sulfur, there are 16 protons and 16 \_\_\_\_\_.  
 27. number of electrons in neutral atom of lithium  
 28. positive charged atom

29. smallest particle of an element that still retains properties of the element  
 30. most of an atoms mass is found here  
 33. negative charged atom  
 35. 33 electrons in neutral atom  
 36. whole number listed on the periodic table  
 37. an atom with the same atomic number and different \_\_\_\_\_

## Down

1. positive subatomic particle  
 2. atoms that lose electrons have a \_\_\_\_\_ charge  
 3. location of protons and neutrons  
 4. atomic number of 28  
 5. neutral subatomic particle  
 7. having no charge  
 8. How many neutrons are in N-16

9. Atoms that gain electrons  
 12. total protons + neutrons  
 15. located in the electron cloud  
 17. To be neutral, protons must equal \_\_\_\_\_.  
 18. charge atom  
 22. center of atom  
 23. potassium has a charge of +1. It has \_\_\_\_\_ electrons  
 24. Cl-35 has 18 \_\_\_\_\_.  
 25. atoms of same element with different masses  
 31. an atom has 11 protons + 12 neutrons =  
 32. number of protons in carbon  
 34. Mg-24 and Mg-25 are \_\_\_\_\_ of the same element