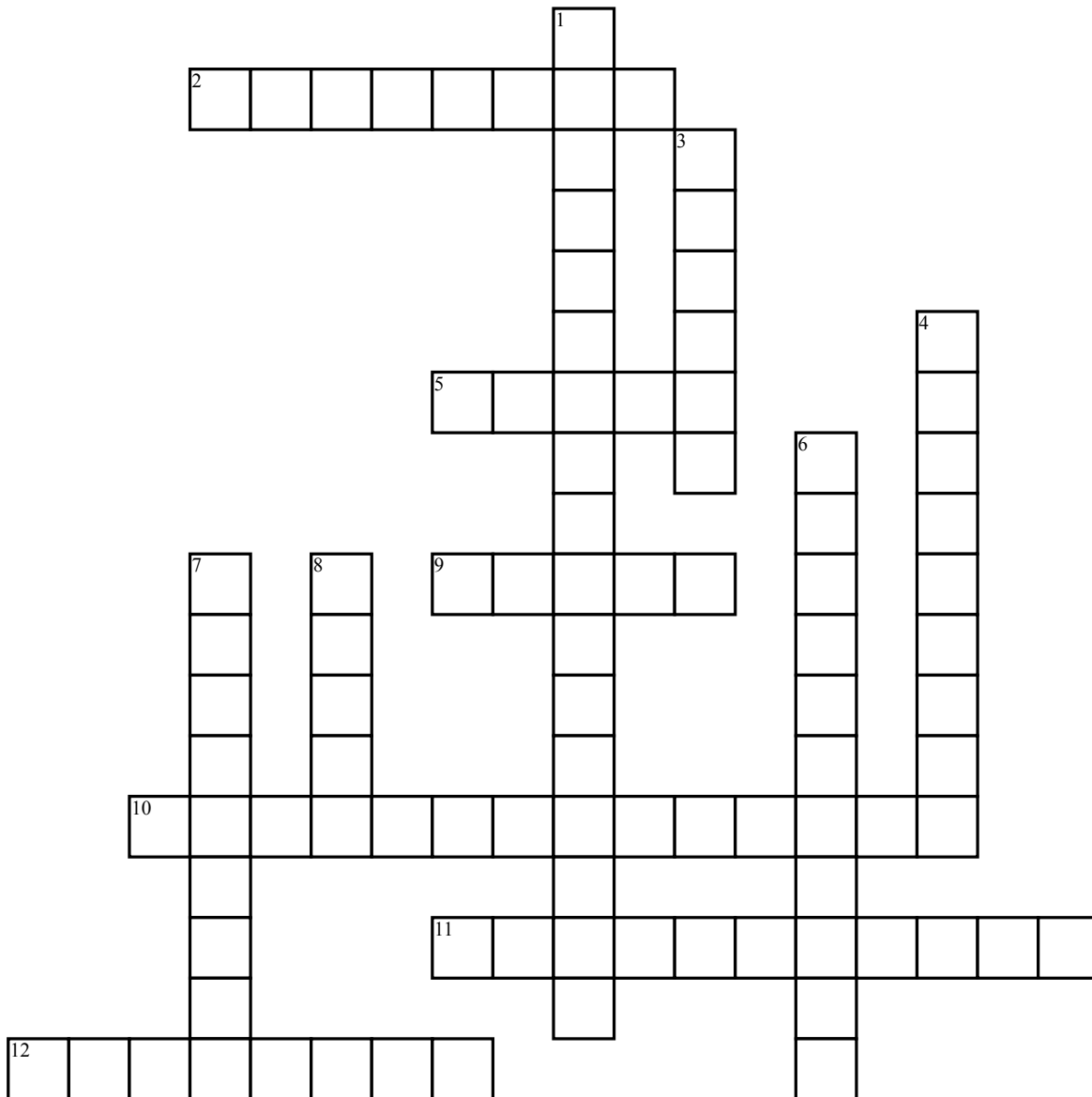


# Acid/Base Physiology



## **Across**

**2.** The Henderson-Hasselbalch equation is used to measure (blank) blood pH

**5.** Phosphate system helps buffer intracellular fluid and...

**9.** The metabolic component involved in maintaining blood pH is from this system

**10.** Law of Mass Action states that reversible reactions will be driven from one side with higher (blank) to the side with lower (blank)

**11.** Metabolic acidosis is caused due to the accumulation of acids not derived from the (blank) system

**12.** If your blood pH is below 7.35, it is in a state of...

## **Down**

**1.** What enzyme catalyzes the reaction gone over in lab?

**3.** Resists changes in pH of a solution

**4.** Respiratory (blank) is when there is little CO<sub>2</sub> in the plasma

**6.** This system is important in buffering blood plasma

**7.** What was the pH indicator used in lab?

**8.** You have a beaker with a solution that raises the pH. What would you call that solution?