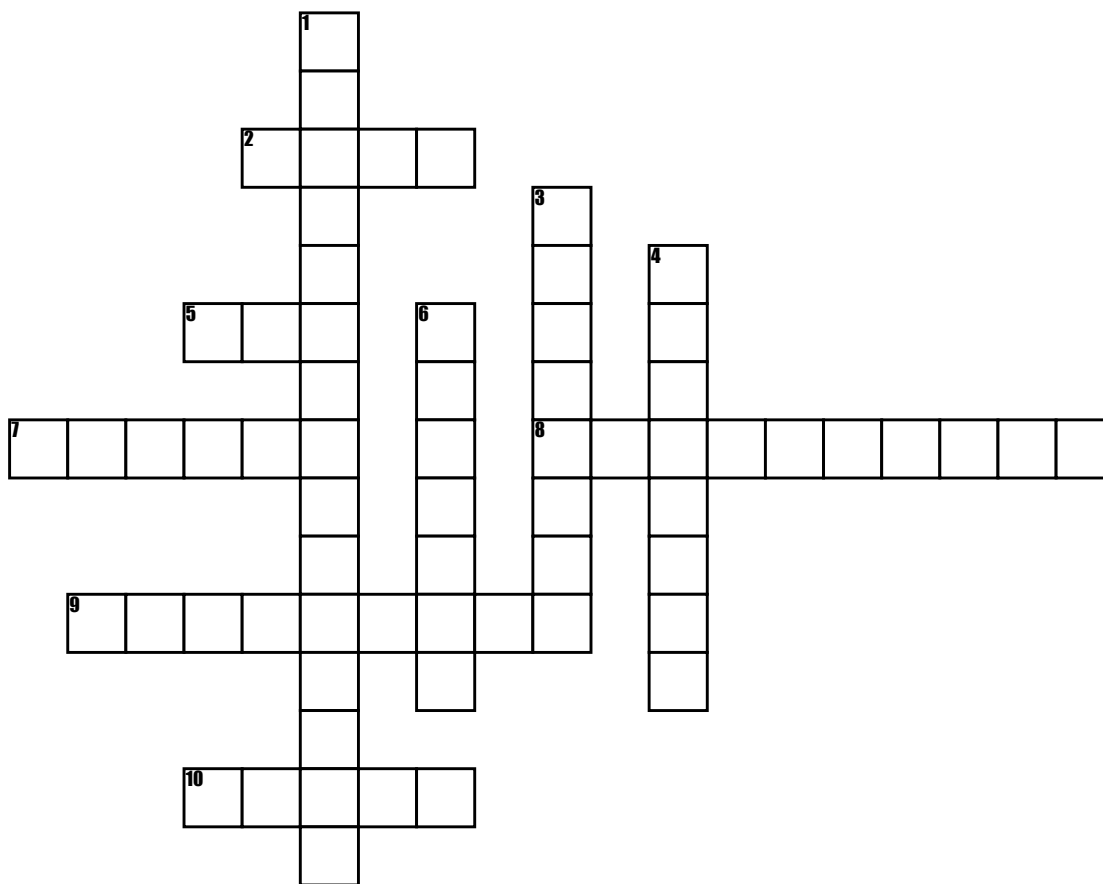


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## Atomic structure, Isotopes, Atomic Mass, Ions, And Atomic Theory



### **Across**

**2.** In 1913, \_\_\_\_\_ proposed a theory for the hydrogen atom based on quantum theory.

**5.** An \_\_\_\_\_ is a charged atom or molecule.

**7.** \_\_\_\_\_ atomic theory proposed that all matter was composed of atoms, indivisible and indestructible building blocks.

**8.** \_\_\_\_\_ overturned Thomson's model in 1911 with his well-known gold foil experiment in which he demonstrated that the atom has a tiny, heavy nucleus.

**9.** \_\_\_\_\_ are found in shells or orbitals that surround the nucleus of an atom.

**10.** \_\_\_\_\_ bond

### **Down**

**1.** \_\_\_\_\_ and \_\_\_\_\_ are found in the nucleus.

**3.** In an isotope, which part of the atom changes?

**4.** \_\_\_\_\_ are variants of a particular chemical element which differ in neutron number.

**6.** \_\_\_\_\_ who discovered the electron in 1897, proposed the plum pudding model of the atom in 1904 before the discovery of the atomic nucleus.