

Name: _____ Date: _____ Period: _____

Atom's

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| 1. is the smallest constituent unit of ordinary matter that has the properties of a chemical element. | A. periods |
| 2. is a subatomic particle, symbol p or p ⁺ , with a positive electric charge of +1e elementary charge and mass slightly less than that of a neutron. | B. atomic mass |
| 3. is a subatomic particle, symbol n or n ⁰ , with no net electric charge and a mass slightly larger than that of a proton. | C. alkali metals |
| 4. is a subatomic particle, symbol e ⁻ or β ⁻ , with a negative elementary electric charge. | D. valence |
| 5. the central and most important part of an object, movement, or group, forming the basis for its activity and growth | E. periodic law |
| 6. an atom or molecule with a net electric charge due to the loss or gain of one or more electrons. | F. protons |
| 7. a positively charged ion, i.e., one that would be attracted to the cathode in electrolysis | G. atomic number |
| 8. a negatively charged ion, i.e., one that would be attracted to the anode in electrolysis. | H. periodic table |
| 9. a negatively charged ion, i.e., one that would be attracted to the anode in electrolysis. | I. nucleus |
| 10. of a chemical element (also known as its proton number) is the number of protons found in the nucleus of an atom of that element, and therefore identical to the charge number of the nucleus. | J. groups |
| 11. is the mass of an atomic particle, sub-atomic particle, or molecule. | K. Ions |
| 12. is an electron that is associated with an atom, and that can participate in the formation of a chemical bond | L. metals |
| 13. a solid material that is typically hard, shiny, malleable, fusible, and ductile, with good electrical and thermal conductivity | M. cations |
| 14. an element or substance that is not a metal. | N. electrons |
| 15. have the same number of atomic orbitals. For example, every element in the top row | O. alkaline earth metals |

16. that the properties of the elements are periodic functions of their atomic numbers	P. noble gases
17. also known as a family	Q. Isotopes
18. horizontal row of the periodic table	R. halogens
19. elements lithium, sodium, potassium, rubidium, cesium, and francium, occupying Group IA (1) of the periodic table.	S. Anions
20. elements beryllium, magnesium, calcium, strontium, barium, and radium, occupying Group IIA (2) of the periodic table.	T. neutrons
21. elements fluorine, chlorine, bromine, iodine, and astatine, occupying group VIIA (17) of the periodic table.	U. atoms
22. elements helium, neon, argon, krypton, xenon, and radon, occupying Group 0 (18) of the periodic table.	V. nonmetals