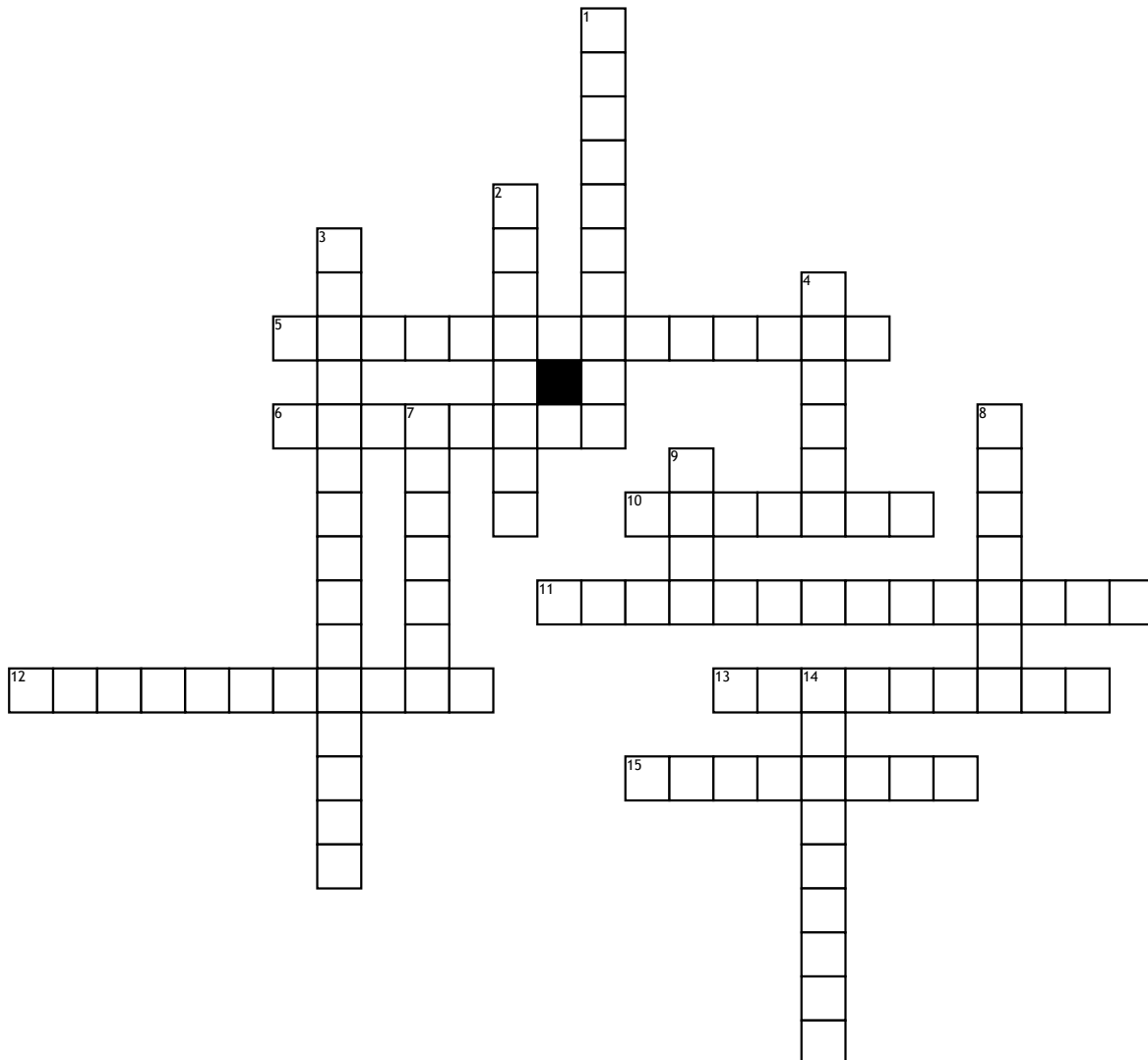


BSMT Final exam



Across

5. Involves the reaction of acid (e.g. HCl) and a base (NaOH) to form salt and water

6. Relates the mass of solvent to the quantity of solute that is dissolved in it (moles of solute / kilograms of solvent)

10. The dispersing medium in a solution.

11. A solution that contains more solute than would dissolve in a saturated solution at a given temperature.

12. A solution in which more solute can be dissolved at a given temperature.

13. A solution that contains the maximum amount of solute that dissolves at a given temperature.

15. A homogeneous mixture of two or more pure substances

Down

1. The maximum quantity of the substance, expressed in grams, that will dissolve in 100 g of solvent at a specific temperature.

2. Relates the volume of solution to the quantity of solute that it contains (moles of solute / liter of solution).

3. The chemical indicator used in acid-base titration.

4. The part of a solution that is being dissolved .

7. A solution whose concentration is to be determined and is placed in an erlenmeyer flask.

8. A solution of known concentration placed in a burette.

9. Obtained by dividing the mass of solute by its molar mass.

14. Used to determine the concentration of an acid or base in a solution.