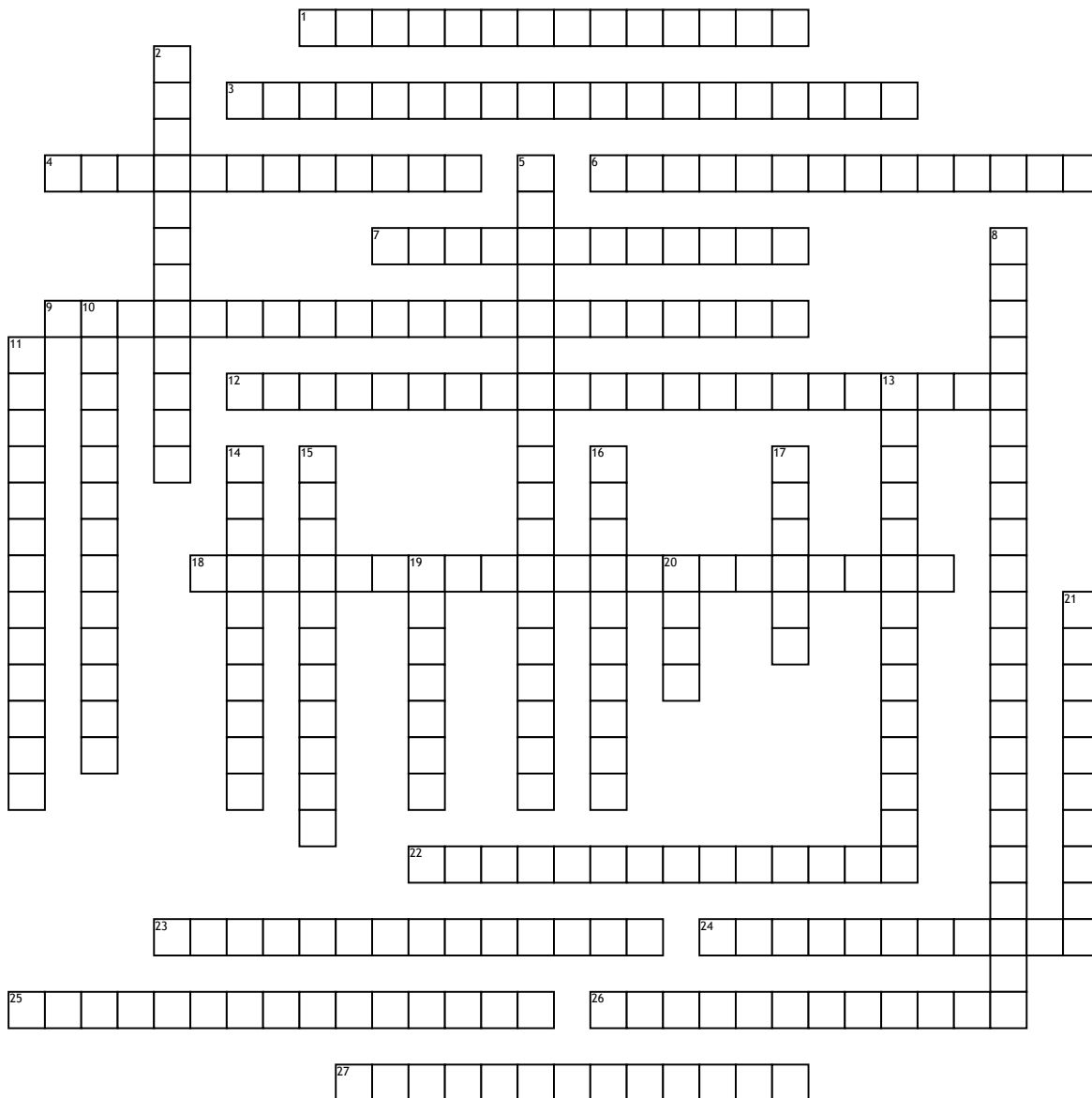


Bases And Acids Part 1



Across

1. A reaction when salt completely dissociates in water, and its anion or cation react with the water to produce hydroxide ions or hydronium ions that affect the pH of the solution.
3. A substance which can behave as either a B/L acid or a B/L base, depending on the circumstances. Water is the prototypical amphoteric substance.
4. H_3O^+ (can be used interchangeably with H^+)
6. Low pOH and high pH
7. A polyprotic acid that has two acidic H^+ ions. An example is H_2SO_4 .
9. Produce H^+ (as H_3O^+) ions in water (the hydronium ion is a hydrogen ion attached to a water molecule) Taste sour Corrode metals Electrolytes React with bases to form a salt and water pH is less than 7 Turns blue litmus paper to red
12. Two substances related to each other by the donating and accepting of a single H^+ ion.

18. Produce OH^- ions in water Taste bitter, chalky Are electrolytes Feel soapy, slippery React with acids to form salts and water pH greater than 7 Turns red litmus paper to blue
22. An acid that has only one acidic H^+ ion.
23. Acid contains H and dissociates to produce H^+ ions in aqueous solution, while a base contains OH and dissociates to produce OH^- ions in aqueous solution.
24. H^+
25. Low pH and high pOH
26. When acids and bases ionize - fall apart - in solution to form electrolyte solutions
27. A polyprotic acid that has three acidic H^+ ions. An example is H_3PO_4 .

Down

2. OH^-
5. An acid is defined as a hydrogen-ion donor and a base is a hydrogen-ion acceptor.
8. A reaction in which an acid and a base in an aqueous solution react to produce a salt and water.

10. The species produced when an acid donates a hydrogen ion to form a base.
11. The species produced when a base accepts a hydrogen ion to form an acid.
13. An acid that has two or more acidic H^+ ions.
14. An acid which contains hydrogen and one other element. Does not contain oxygen.
15. $\text{pOH} = -\log[\text{OH}^-]$
16. Chemicals that change color in the presence of acids or bases.
17. An indicator that is used to determine if a solution is acidic or basic. Red litmus turns blue for bases, while blue litmus turns red for acids.
19. Any acid that contains hydrogen and an oxyanion.
20. An ionic compound made from the cation from a base, and an anion from an acid
21. $\text{pH} = -\log[\text{H}^+]$