Name: _____ Date: _____

CHM-113 Matching Terms

1. Visible light that is emitted appears as colored lines corresponding to its	A. momentum
2. Emissions	B. light
3. Visible light that is absorbed appears as dark lines corresponding to its	C. orbital
4. Absorption	D. state
5. Release of energy as an electron moves from a higher energy level to a lower energy	E. number
6. Gain of energy as an electron moves from a lower energy level to a higher energy	F. n
7. Emission and transmission of energy in the form of waves that move through space at the speed of	G. state
8. Electromagnetic	H. wavelength
9. Length of a wave measured from peak to peak or trough to trough in units of	I. level
10. Cycles per cycle of a wave in units of s-1 or	J. spectra
11. Packet of energy in the form of	K. absorbed
12. Specified quantity of energy that can be emitted or	L. level
13. Light behaves as both a wave and a particle and requires a quanta of energy to	M. m
14. Photoelectric	N. electron
15. Electrons in a higher energy level than their lowest energy, ground	O. emit
16. Excited	P. effect
17. Electrons in their lowest energy	Q. numbers
18. Ground	R. Radiation
19. Set of four numbers that describe the location and spin of an	S. l
20. Quantum	T. Hz
21. Whole number (non-zero) that describes the energy level of an orbital, represented by	U. n
22. Principle quantum	V. state
23. Energy level of an electron, represented by	W. level
24. Whole number (includes 0) that describes the shape of an orbital, represented by	X. spectra

25. Angular

26. Represented by nl that describes the energy level and shape of an

Y. wavelength

Z. light