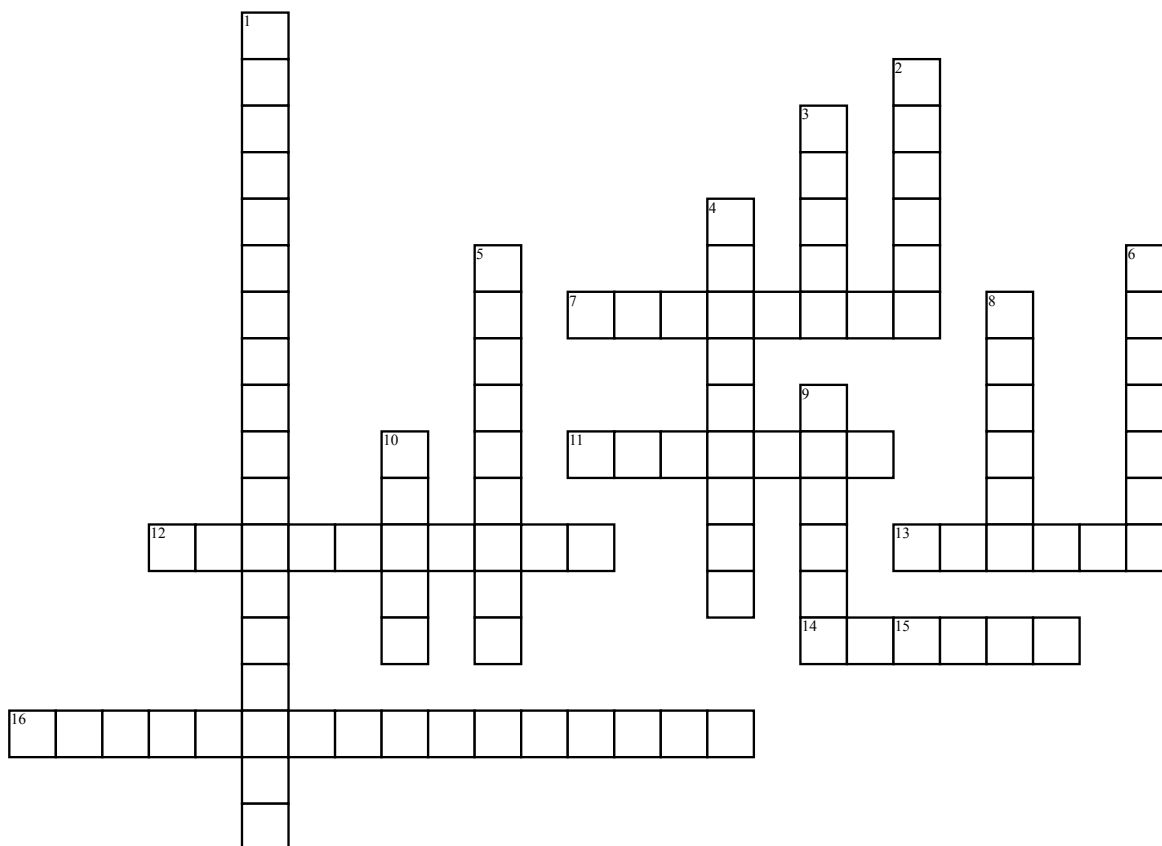# Chaper 10 Review



**Across**

**7.** Find the empirical formula of 0.463g Ti, 0.0544g of C , 0.00685 of H, 0.0725 g O

**11.** How many grams are in 235 moles arsenic

**12.** the mass of an atom of a chemical element expessed in__ __ units

**13.** Ho many moles are in 22 grams of Argon

**14.** How many moles are in 23 grams of gold

**16.** a formula giving the proportions of the elements present in a compound but not the actual numbers or arrangment of atoms

**Down**

**1.** the percentage by mass of each element in a compond

**2.** How many moles are in 10 grams of Hydrogen

**3.** Find the empirical formula of 0.0115g of O, 0.0134g Iron, 0.00759g S

**4.** How many grams are in 13 moles of chromium

**5.** physical property defined as the mass of a given substance._ are almost always expressed in g/mol

**6.** How many grams are in 2 moles of sulfur

**8.** How many moles are in 45,1 grams of magnesium

**9.** how many moles are in 15 grams of lithium

**10.** Empirical Formula of 32.8% Chromium and 67.2% Chlorine

**15.** Agravo's number is how many moles?