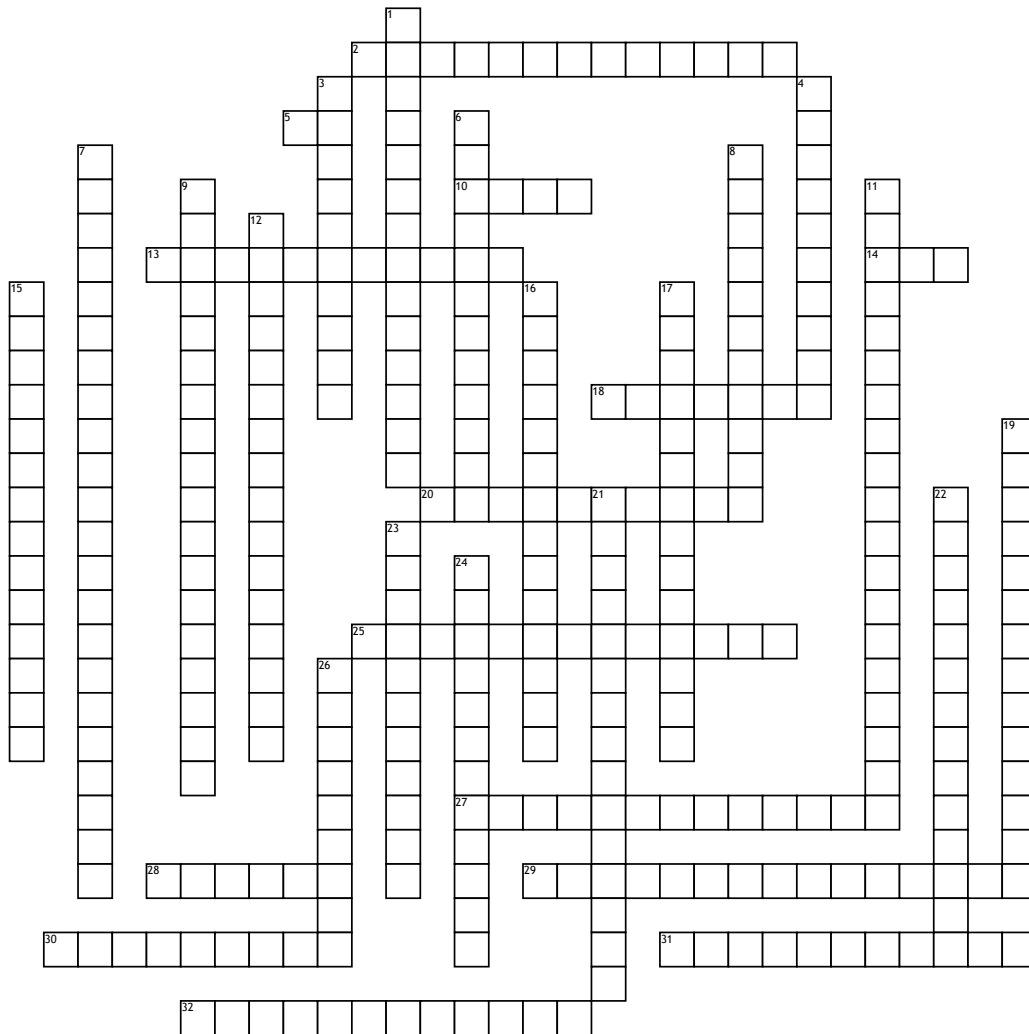


Chapter 19 Crossword



Across

2. The species produced when a base accepts a hydrogen ion to form an acid.
 5. A measure of the strength of an acid or base solution which is based on the amount of H^+ ion.
 10. An ionic compound made from the cation from a base, and an anion from an acid.
 13. $pOH = -\log[OH^-]$
 14. A measure of the strength of an acid or base solution which is based on the amount of OH^- ion.
 18. Any acid that contains hydrogen and an oxyanion.
 20. Chemicals that change color in the presence of acids or bases.
 25. A polyprotic acid that has three acidic H^+ ions. An example is H_3PO_4 .
 27. The species produced when an acid donates a hydrogen ion to form a base.
 28. An indicator that is used to determine if a solution is acidic or basic. Red litmus turns blue for bases, while blue litmus turns red for acids.
 29. Low pH and high pOH
 30. Acids that only ionize partially in solution.

Word Bank

Monoprotic Acid
 Polyprotic Acid
 Weak Bases
 Dissociation
 Acidic Solutions
 Oxyacid
 Hydronium ion

Basic Solutions
 Diprotic Acid
 pOH Equation
 Bronsted-Lowry Model
 Basic solutions
 Indicators
 Binary Acid

Neutralization Reaction
 Weak Acids
 Strong Bases
 pH Equation
 Strong Acids
 Salt

pOH
 Triprotic Acid
 Amphoteric Substance
 Acidic solutions
 Conjugate acid
 Arrhenius Model

pH
 Salt Hydrolysis
 Hydrogen ion
 Litmus
 Neutral solutions
 Conjugate Base

31. Bases that dissociate entirely into metal ions and hydroxide (OH^-) ions in aqueous solution (Arrhenius base).
 32. H_3O^+ (can be used interchangeably with H^+)

Down

1. An acid that has only one acidic H^+ ion.
 3. $pH = -\log[H^+]$
 4. An acid which contains hydrogen and one other element. Does not contain oxygen.
 6. When acids and bases ionize - fall apart - in solution to form electrolyte solutions.
 7. A reaction in which an acid and a base in an aqueous solution react to produce a salt and water.
 8. Acids that ionize completely in solution.
 9. An acid is defined as a hydrogen-ion donor and a base is a hydrogen-ion acceptor.
 11. A substance which can behave as either a B/L acid or a B/L base, depending on the circumstances.
 12. Have $pH = 7$
 15. A reaction when salt completely dissociates in water, and its anion or cation react with the water to produce hydroxide ions or hydronium ions that affect the pH of the solution.

16. Acid contains H and dissociates to produce H^+ ions in aqueous solution, while a base contains OH and dissociates to produce OH^- ions in aqueous solution.

17. An acid that has two or more acidic H^+ ions.
 19. Have $pH > 7$
 21. Have $pH < 7$
 22. Low pOH and high pH
 23. H^+
 24. A polyprotic acid that has two acidic H^+ ions. An example is H_2SO_4 .
 26. Bases that ionize only partially in dilute aqueous solution to form the conjugate acid and hydroxide ions.