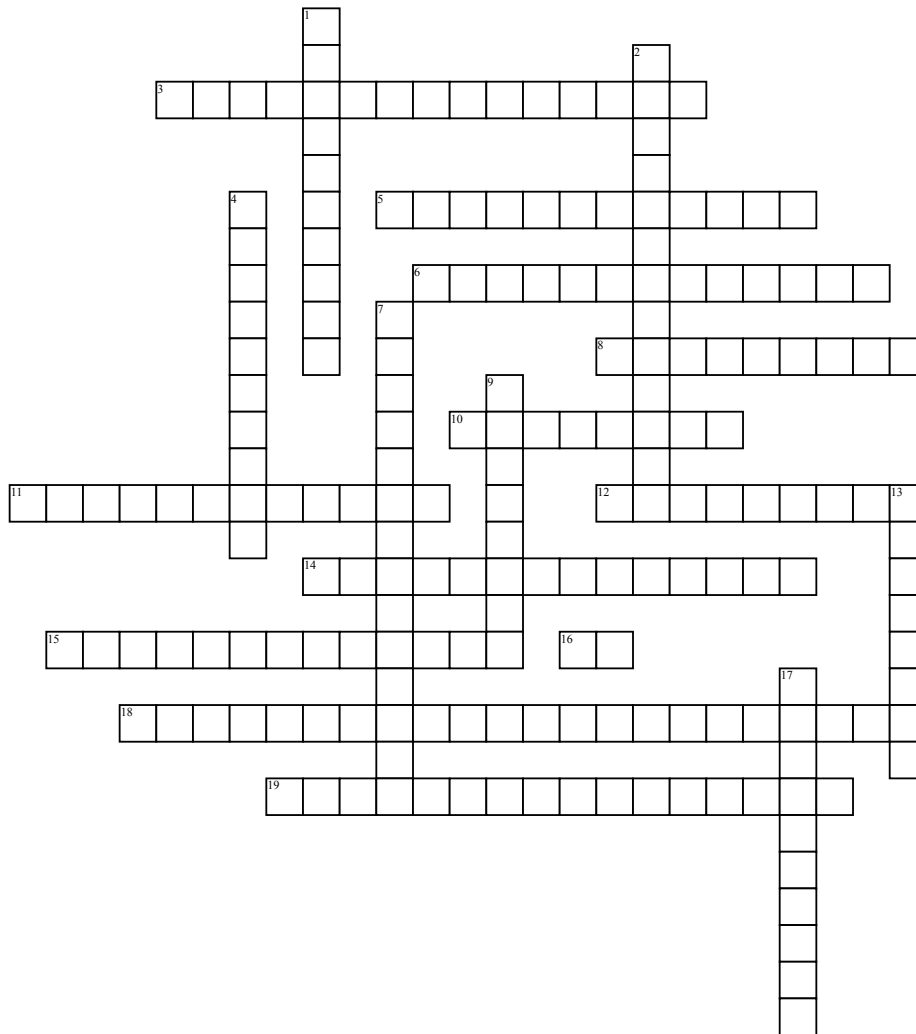


Chapter 19



Across

3. An aqueous solution in which the concentration of hydrogen and hydroxide ions are equal
 5. Any solution in which the hydroxide ion concentration is greater than the hydrogen ion concentration
 6. Two substances that are related by the loss or gain of a single hydrogen ion
 8. Any Substance that can accept a pair of electrons to form a covalent bond.
 10. A base that reacts with water to form the hydroxide ion and the conjugate acid of the base
 11. (H₃O⁺) The positive ion formed when a water molecule gains a hydrogen ion.

12. Any substance that can donate a pair of electrons to form a covalent bond.

14. Any solution in which the hydrogen ion concentration is greater than the hydroxide ion concentration.

15. The particle formed when a base gains a hydrogen ion

16. A number used to denote the hydrogen-ion concentration, or acidity, of a solution

18. A reaction in which an acid and a base react in an aqueous solution to produce a salt and water

19. A solution of known concentration used in carrying out a titration

Down

1. An acid that is completely (or almost completely) ionized in aqueous solution.
 2. is the particle that remains when an acid has donated a hydrogen ion.
 4. A substance that can act as both an acid and a base.
 7. A term describing the reaction in which two water molecules react to produce ions
 9. An acid that is only slightly ionized in aqueous solution
 13. The point in a titration at which the indicator changes color
 17. A base that completely dissociates into metal ions and hydroxide ions in aqueous solution.

Word Bank

Self-Ionization
 Base solution
 Strong Base
 conjugate base
 Neutral Solution

Strong Acid
 conjugate acid
 Acidic solution
 Lewis acid
 Lewis Base

Amphoteric
 Weak Acid
 standard solution
 conjugate pair
 Neutralization Reaction

pH
 Hydronium ion
 End point
 weak base