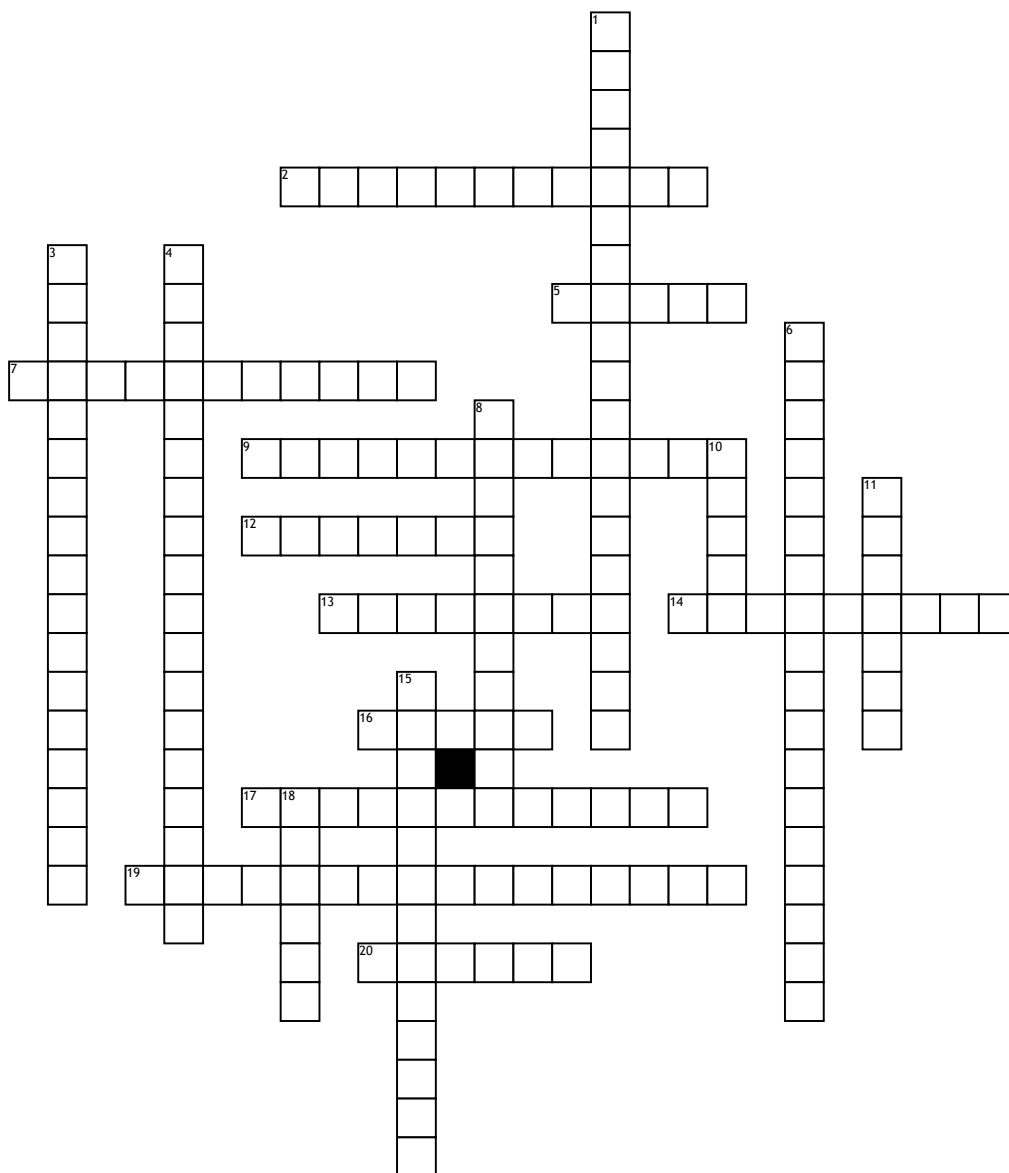


Name: _____

Date: _____

Chapter 3 Chemistry



Across

2. A SI unit of measurement comprised of a combination of the seven base units.
5. JOULE basic SI unit of energy
7. A measure of the "hotness" or "coldness" of a substance.
9. is substance accepted by almost all scientists
12. meter, kilogram, second, ampere, kelvin, mole, candela
13. refers to the closeness of a measured value to a standard or known value.
14. refers to how close two or more measurements are to each other, regardless of whether those measurements are accurate or not.

16. the volume of one kilogram of pure water.

17. is a measure of how inaccurate a measurement is, standardized to how large the measurement is.

19. is a ratio (or fraction) which represents the relationship between two different units.

20. is the name of the force exerted on an object due to the acceleration of gravity.

Down

1. a way chemists and other scientists convert units of measurement.

3. denotes the value of a substance's properties found in a localized lab

4. A method of expressing numbers in terms of a decimal number between 1 and 10 multiplied by a power of 10.

6. a number are digits that carry meaning contributing to its measurement resolution.

8. collection of quantitative or numerical data that describes a property of an object or event.

10. is the discrepancy between a computed, observed, or measured value or condition and the true, specified, or theoretically correct value or condition.

11. is defined as the ratio between mass and volume or mass per unit volume.

15. energy an object possesses due to its motion.

18. defined in science as the ability to do work.