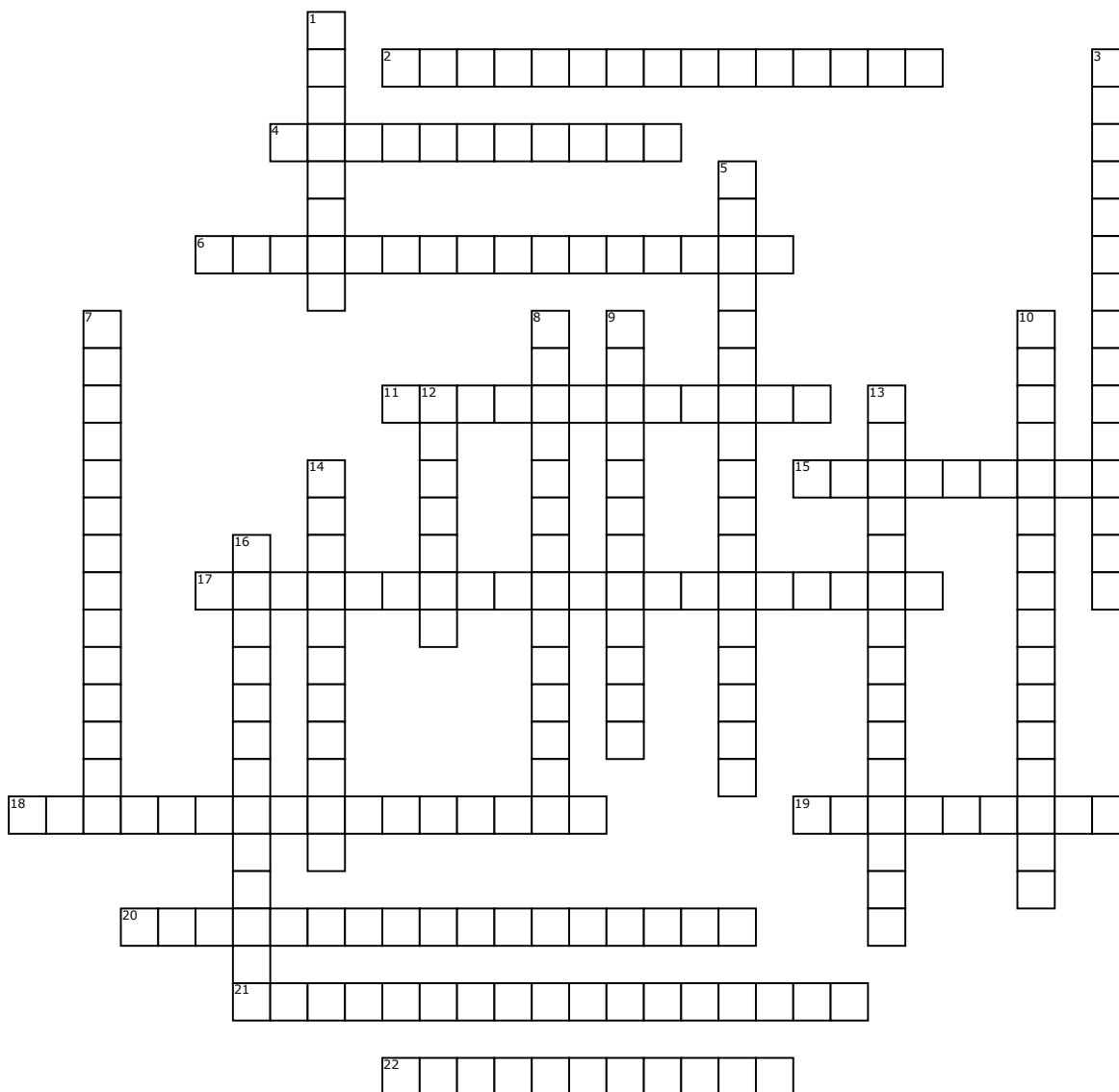


Chem 20 part 1



Across

2. Electrons in the outer shell of an atom.

4. The atom that is the most electronegative.

6. A formula giving the proportions of the elements present in a compound but not the actual numbers or arrangement of atoms.

11. A chemical bond that involves the sharing of electron pairs between atoms.

15. The complete transfer of valence electron(s) between atoms.

17. Chemical bond where two atoms share a pair of electrons with each other.

18. A formula giving the number of atoms of each of the elements present in one molecule of a specific compound.

19. Every atom wants to have eight valence electrons in its outermost electron shell.

20. A formula that shows the arrangement of atoms in the molecule of a compound.

21. A measure of the tendency of an atom to attract a bonding pair of electrons.

22. A structural representation of a molecule where dots are used to show electron position around the atoms and lines or dot pairs represent covalent bonds between atoms.

Down

1. A pair of valence electrons that are not shared with another atom.

3. The electrostatic attraction between the positively charged atomic nuclei of metal atoms

5. Bond between two atoms in which electrons are shared unequally.

7. A chemical formula that indicates the spatial arrangement of atoms or atomic groupings in a molecule.

8. The outermost orbital of an atom.

9. The maximum number of bonds an element can have.

10. A molecule in which the electrons are shared equally between the nuclei.

12. An energy state in the atomic model which describes where an electron will likely be.

13. Electrons that are not shared with another atom.

14. a model used to predict the geometry of individual molecules from the number of electron pairs surrounding their central atoms.

16. An asymmetric molecule with non-uniform positive and negative charges.