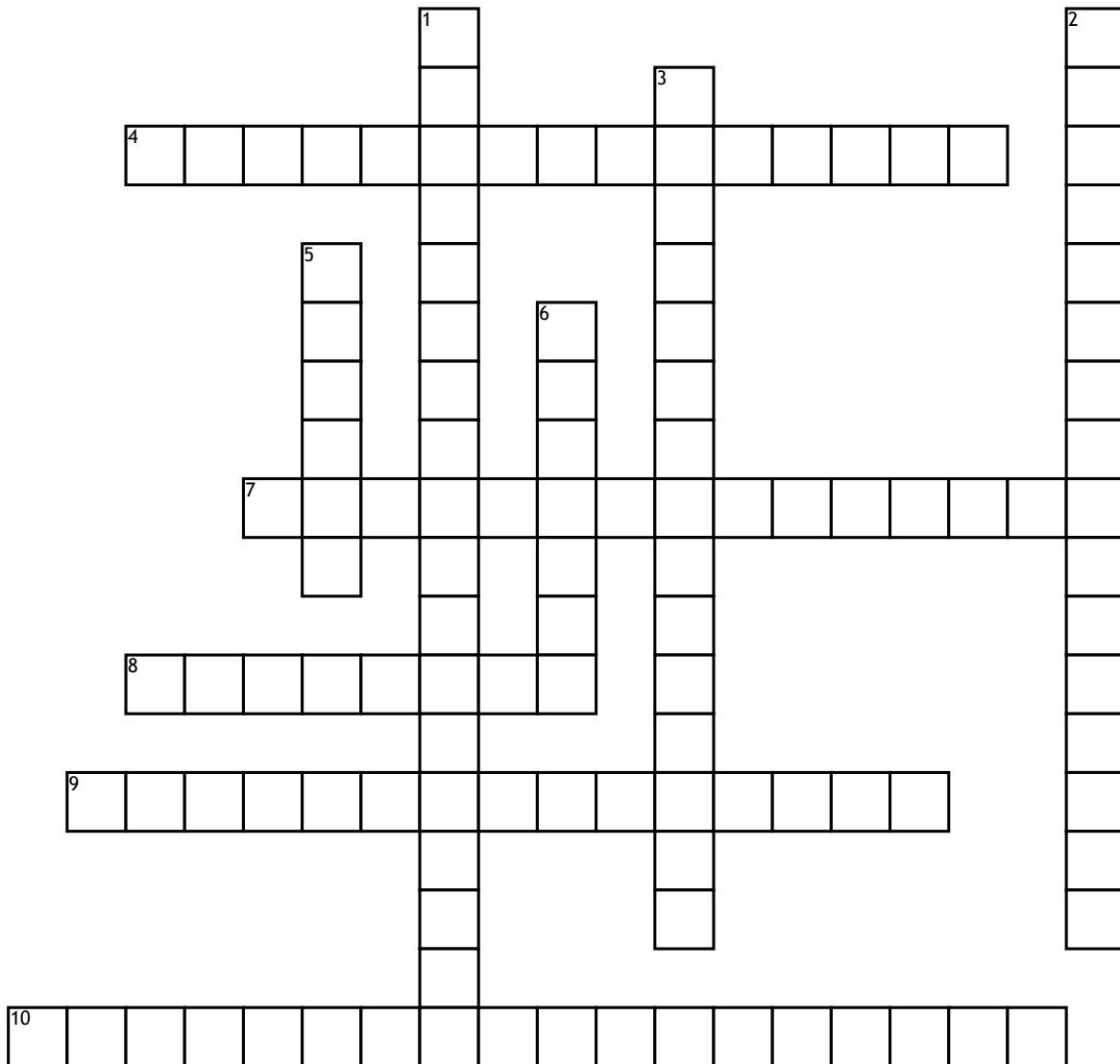


Name: \_\_\_\_\_

Date: \_\_\_\_\_

# Chem Equations



## Across

4.  $F_{el} = k q_1 q_2 / r^2$

7.  $V/T = k$

8. moles of solute / liters of solution

9.  $V = k q_1 q_2 / r$

10.  $M_i V_i = M_f V_f$

## Down

1.  $n \times \text{molar mass of element} / \text{molar mass of compound} \times 100\%$  - where  $n$  = the number of moles of the element in one mole of the compound

2.  $^{\circ}\text{C} = (^{\circ}\text{F} - 32) \times 5/9$

3.  $V = kn$

5.  $\text{actual yield} / \text{theoretical yield} \times 100\%$

6.  $d = m/v$