

Name: \_\_\_\_\_ Date: \_\_\_\_\_

# Chemical Interactions

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|---|--------------------------------|
| 1. Produces new substances by changing the way in which atoms are arranged                                  | A. products                    |
| 2. The substances present at the beginning of a chemical reaction   | B. endothermic reaction        |
| 3. The substances formed by a chemical reaction   | C. base                        |
| 4. When two chemical liquids mix together and form a solid  | D. law of conservation of mass |
| 5. A reaction in which a new compound is formed by the combination of simpler reactants                     | E. decomposition               |
| 6. A reaction in which a reactant breaks down into simpler products   | F. alloy                       |
| 7. A reaction in which one reactant is always oxygen and another reactant often contains carbon or hydrogen | G. respiration                 |
| 8. A substance that increases the rate of a chemical reaction but is not itself consumed in the reaction    | H. solution                    |
| 9. States that in a chemical reaction, atoms are neither created nor destroyed                              | I. exothermic reaction         |
| 10. The number in front of the chemical formula   | J. suspension                  |
| 11. The energy associated to bonds  | K. reactants                   |
| 12. A reaction in which energy is released  | L. photosynthesis              |
| 13. A reaction in which energy is absorbed  | M. solvent                     |
| 14. Plants absorb energy from sunlight  | N. synthesis                   |
| 15. Living cells obtain energy from glucose molecules through the process called                            | O. catalyst                    |
| 16. A type of mixture that is the same throughout   | P. acid                        |
| 17. A substance that is dissolved to make a solution  | Q. dilute                      |
| 18. A substance that dissolves a solute   | R. combustion                  |
| 19. Particles that do not dissolve  | S. solute                      |
| 20. The amount of solute dissolved in a solvent at a particular temperature                                 | T. bond energy                 |

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| 21. A solution that has a low concentration of solute         | U. concentration     |
| 22. The amount of solute that can be dissolved in the solvent | V. coefficient       |
| 23. A substance that has a pH lower than 7                    | W. precipitate       |
| 24. A substance that has a pH higher than 7                   | X. neutral           |
| 25. A substance that has a pH of 7                            | Y. chemical reaction |
| 26. A mixture of a metal and one or more other elements       | Z. solubility        |