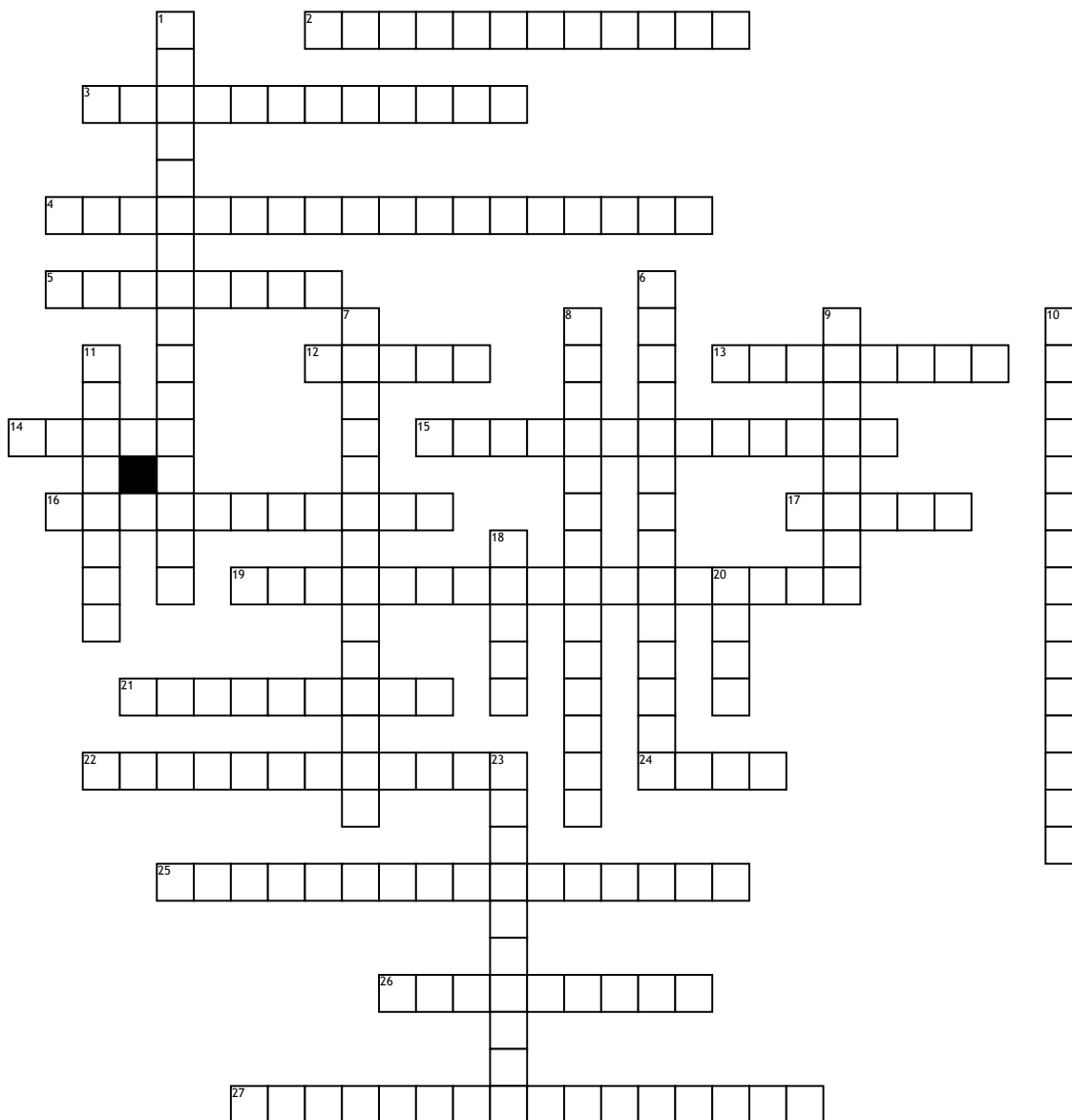


Chemistry



Across

2. An ion that appears on both side of an equation and is not directly involved in the reaction. Ch.11
 3. Small whole numbers that are placed in front of the formulas in an equation in order to balance it. Ch.11
 4. Mass of element divided by mass of compound multiplied by 100. Ch.10
 5. The elements in group 8A of the periodic table. Ch.6
 12. Four Ch.9
 13. A substance that speeds up the reaction but is not used up in the reaction. Ch.11
 14. Five Ch.9
 15. Composed of more than one atom. Ch.9

16. When elements are arranged in order of increasing atomic number, there is a periodic repetition of their physical and chemical properties. Ch.6
 17. Seven Ch.9
 19. $(\text{Mass})(\%)+(\text{mass})(\%)/100$
 21. The mass of a mole of an element. Ch.10
 22. Consist of a single atom with a positive or negative charge resulting from the loss or gain of one or more valence electrons. Ch.9
 24. Ten Ch.9
 25. HCl Ch.9
 26. Generally has properties that are similar to those of metals and nonmetals. Ch.6
 27. Each side of the equation has the same number of atoms of each element and mass is conserved. Ch.11

Down

1. A representation of a chemical reaction; the formulas of the reactants are connected by an arrow with the formulas of the product. Ch.11
 6. HI Ch.9
 7. HClO4 Ch.9
 8. Composed of two elements and can be either ionic or molecular. Ch.9
 9. The no metals of group 7A. Ch.6
 10. Shows the kinds and numbers of atoms in the smallest representative unit of a substance. Ch.7
 11. Poor conductors of heat and electric current. Ch.6
 18. H2O Ch.9
 20. One Ch.9
 23. HNO3 Ch.9