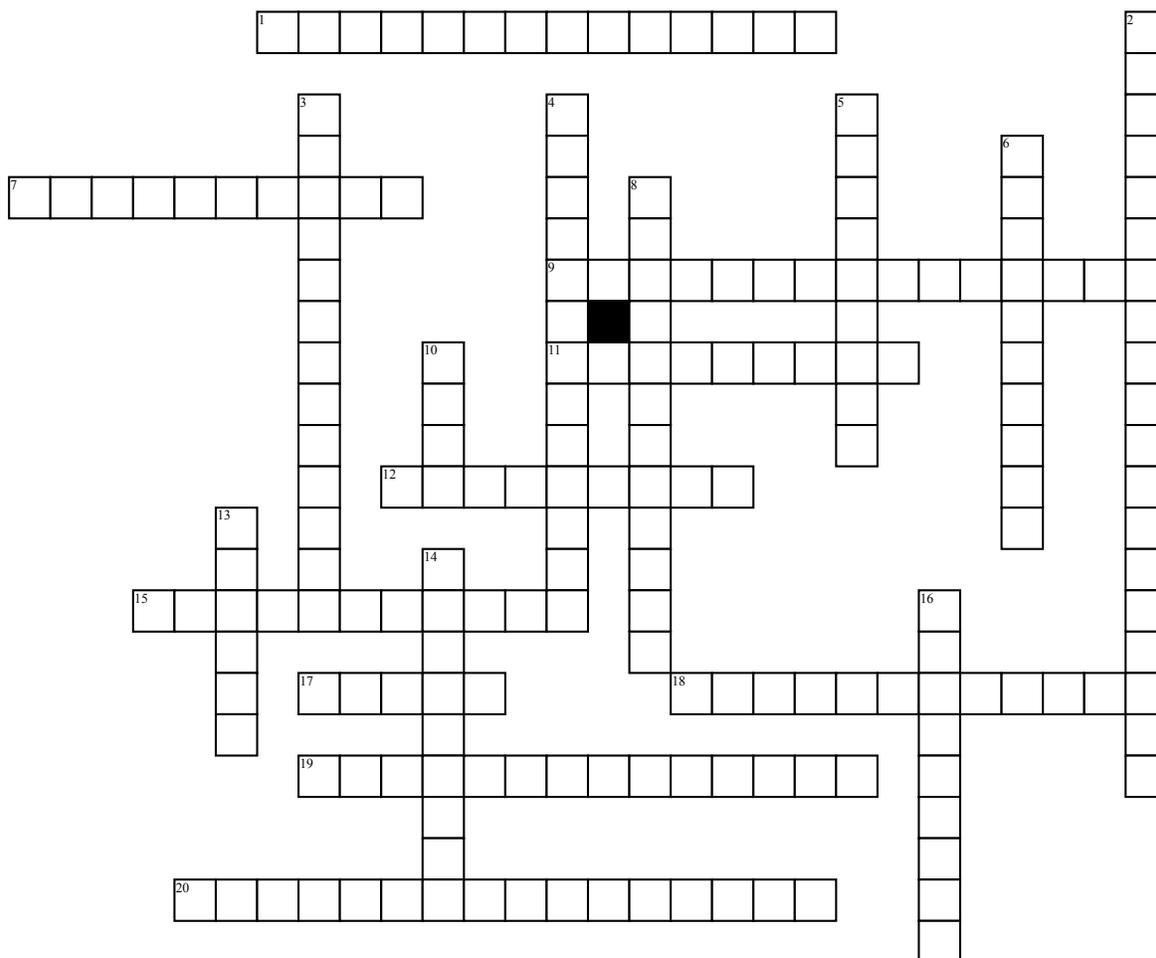


Name: _____

Date: _____

Chemistry 3



Across

1. A unit of mass equal to one-twelfth the mass of a carbon- 12 atom.

7. A hydrocarbon substituent.

9. The number of representative particles contained in one mole of a substance.

11. An organic compound having amino (-NH₂) and carboxyl (-COOH) groups in the same molecule.

12. A substance that act as both an acid and a base.

15. A halocarbon in which one or more halogen atoms are attached to the carbon atoms of an aliphatic chain.

17. Any atom or group of atoms with a negative charge.

18. one-half the distance between the nuclei of two atoms of the same element when the atoms are joined.

19. Describes a solid that lacks an ordered internal structure.

20. A chemical equation in which mass is conserved.

Down

2. Equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

3. A mathematical expression describing the probability of finding an electron at various locations.

4. A positively charged particle emitted from certain radioactive nuclei.

5. Synthesis processes in the metabolism of cells.

6. The weighted average of the masses of the isotopes of an element.

8. The number of protons in the nucleus of an atom of an element.

10. The smallest particle of an element that retains its identity in a chemical reaction.

13. A hydrocarbon containing a carbon-carbon triple bond.

14. One of two or more different molecular forms of an element in the same physical state.

16. An instrument used to measure atmospheric pressure.