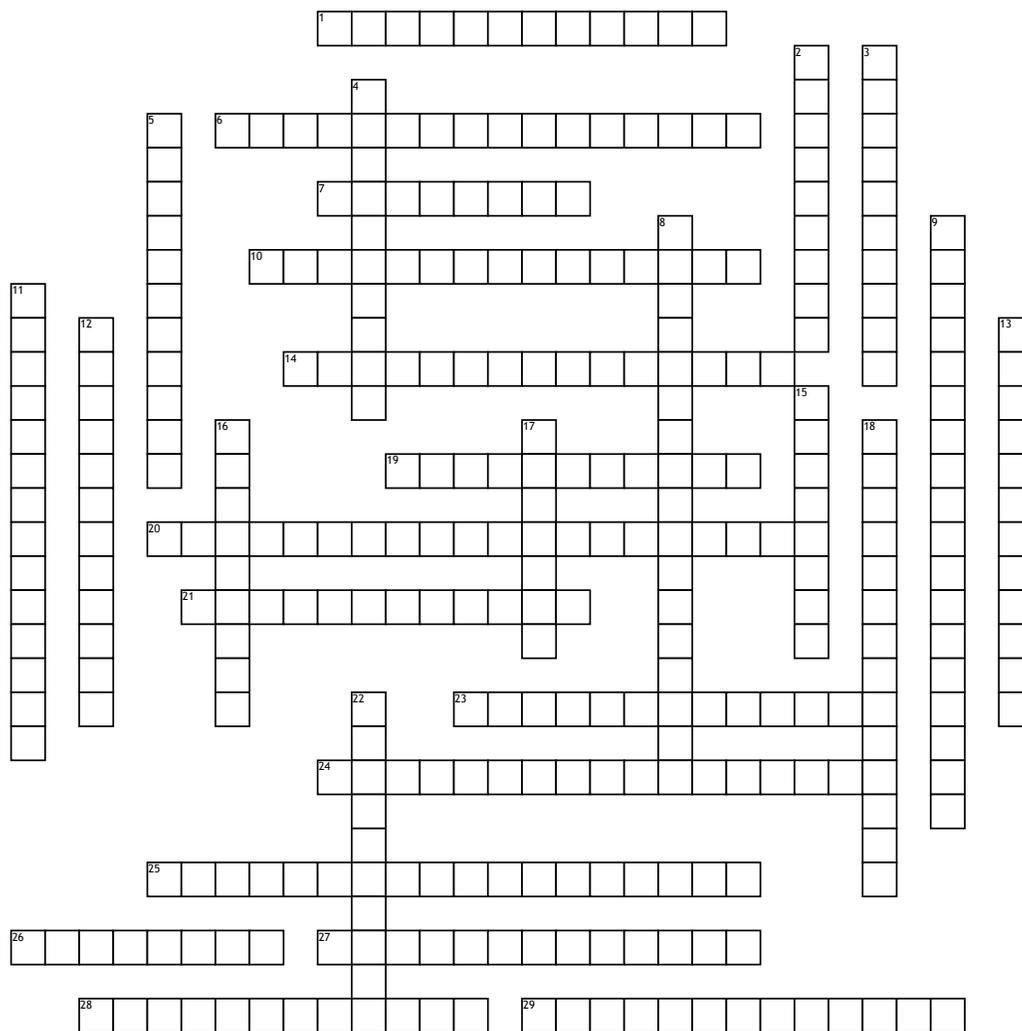


Chemistry Chapter 11 Summary



Across

1. temp at which vapor pressure equals the applied pressure
 6. London Forces, Non Polar dipole-dipole
 7. attraction between particles of the liquid
 10. water rises up narrow capillary tubes easily
 14. no distinct shape or underlying pattern, when split irregular fragments occur
 19. when vaporization occurs in a non-boiling liquid
 20. electrostatic attractions between molecules
 21. temp at which solid turns to liquid
 23. pressure exerted by the evaporated molecules over a liquid
 24. shattered, smaller shapes form with similar edges and angles

25. condensation nullifies the effect of vaporization
 26. intermolecular attractions within the liquid and between the liquid and the container
 27. released when gaseous particles form crystals
 28. reverse of vaporization
 29. one atom is hydrogen and the other is either fluorine, oxygen, or nitrogen

Down

2. liquids ability to resist flowing
 3. different forms of allotropic elements
 4. gas to solid
 5. solid to gas
 8. align themselves so that the positive end of one of the molecule is near the negative end of another molecule

9. pure elements that are polymorphous
 11. crystal's 3-d pattern, depends on number and kinds of particles involved
 12. can form more than one type of crystal lattice
 13. molecules with above-average speeds sometimes break away from the liquid phase
 15. attraction between particles of liquid and particles of other materials
 16. contain the fundamental patterns of the lattices
 17. rapid change between liquid and gaseous states caused by heat
 18. elastic skin that forms on the surface of liquids
 22. break down the normal surface tension of water by interfering with hydrogen bonds

Word Bank

hydrogen bonds
 vaporization
 deposition
 lattice energy
 meniscus
 cohesion

surface tension
 amorphous solids
 allotropes
 allotropic elements
 dispersion forces
 melting point

vapor pressure
 intermolecular forces
 boiling
 adhesion
 capillary action
 evaporation

polymorphous
 crystalline solids
 dynamic equilibrium
 viscosity
 unit cells
 boiling point

crystal lattice
 sublimation
 condensation
 dipole-dipole force
 surfactant