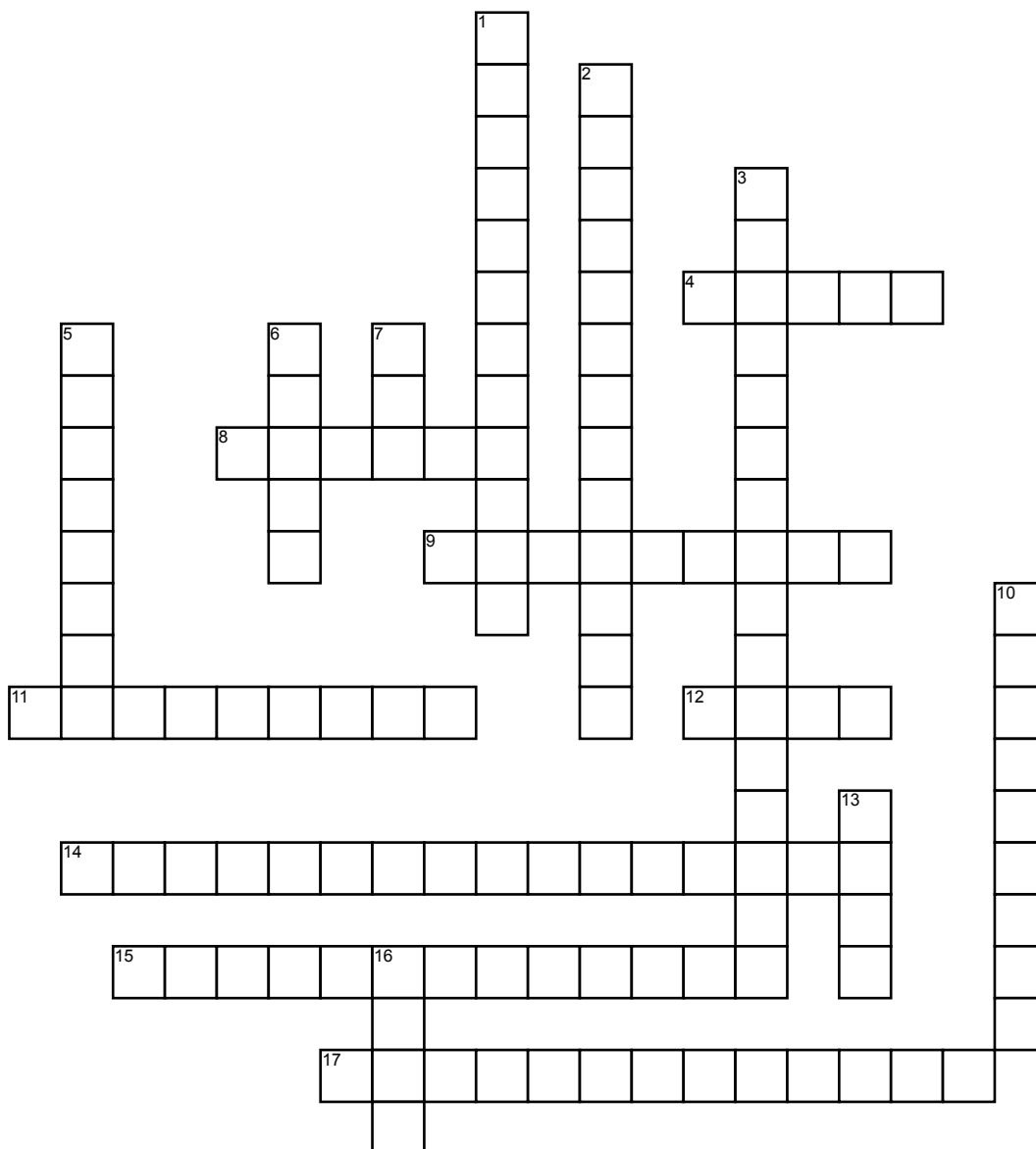


Name: _____ Date: _____ Period: _____

Chemistry Crossword by Brianna Bowen



Across

4. Make the empirical formula for a compound that is 71.65% Cl, 24.27% C, and 4.07% H

8. Make the empirical formula for a compound that is 60.0% C, 4.5% H, and 35.5% O

9. The mass in grams of one mole of any pure substance.

11. Convert 5,000 grams K in moles

12. 6.02×10^{23} is the same thing as saying it's a ____.

14. Also known as the "e formula"; a formula that shows the smallest whole-number mole ratio of the elements of a compound, and might or might not be the same as the actual molecular formula.

15. Convert 321 molecules of Ca into grams

17. Convert 29.54 grams to atoms

Down

1. Convert 270 molecules into moles

2. Convert 3.65 moles Mg to molecules

3. A formula that specifies the actual amount of atoms of each element in one molecule of a substance.

5. How many individual atoms are in a sample of carbon that weighs 7.89 grams?

6. Find the molar mass of the compound Sodium Carbonate (Na_2CO_3)

7. Make the empirical formula for a compound composed of 2.4 grams C and 0.8 grams of hydrogen

10. Convert 10.0 moles Na in grams

13. A measure that reflects the amount of matter.

16. The molar mass of Nitrogen