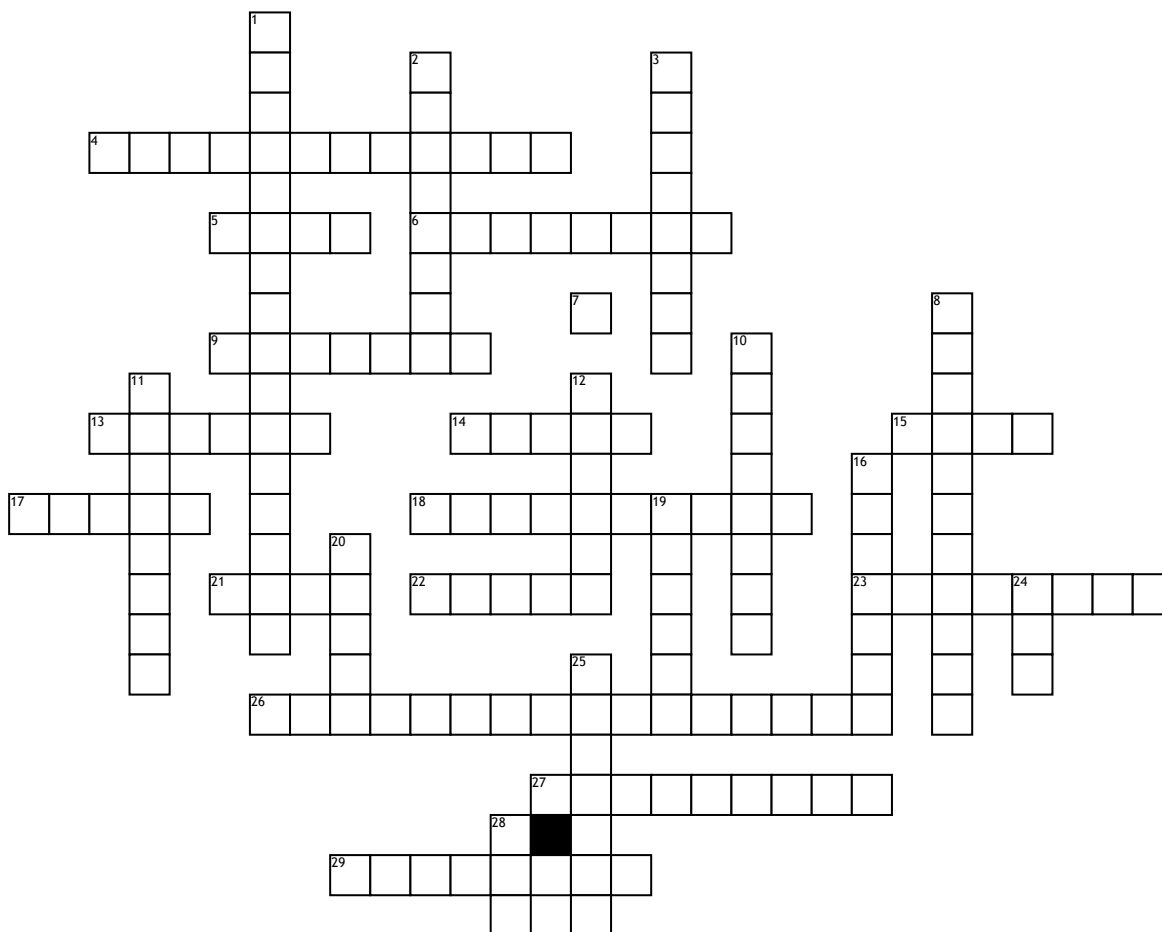


Name: _____

Date: _____

Chemistry Review



Across

4. Gas mingles easily with another porous substance is an example of _____.
5. To solidify a gas, you need a low temperature and a _____ pressure.
6. Thermochemistry studies the _____ of energy.
9. The larger component of the solution is the _____.
13. _____ gas law focuses on the relationship between pressure and volume.
14. _____ volume is the volume occupied by Avogadro's number of particles.
15. A pH value of 9 indicates a weak _____.
17. A _____ gas behaves according to the kinetic molecular theory.
18. The inter-molecular force that acts on non-polar molecules is called _____.
21. _____ laws are determined experimentally.

22. If the reaction rate doubles when the concentration of one reactant doubles, the reaction is the _____ order.
23. A long arrow pointed to the left indicates that a reaction is _____ favored.
26. _____ is the minimum amount of energy needed before a reaction can begin.
27. If the solubility of a solution is 28.2 g/100 mL, the term that describes a 100 ml solution with 28.2 g of solute is _____.
29. _____ bonds occur only with fluorine, oxygen, and nitrogen.

Down

1. The _____ theory states that particles are in constant, random motion.
2. Together the solute and the solvent make a _____.
3. For a constant quantity of gas at a constant temperature, if you increase the volume, the pressure will _____.
7. The abbreviation for molarity is _____.

8. If the free energy value is positive, the reaction is _____.
10. The Tyndall effect is used to distinguish _____ from solutions.
11. _____ is described as moles of solute per liter.
12. _____ heat changes the state but not the temperature of a substance.
16. _____ is the measure of energy dispersal.
19. The smaller component of the solution is the _____.
20. Thermodynamics answers the question, "Can two substances _____ together?"
24. In a heterogeneous reaction, the reactant exist in at least _____ phases.
25. A Bronsted-Lowry acid _____ a proton.
28. $\text{pH} + \text{_____} = 14$