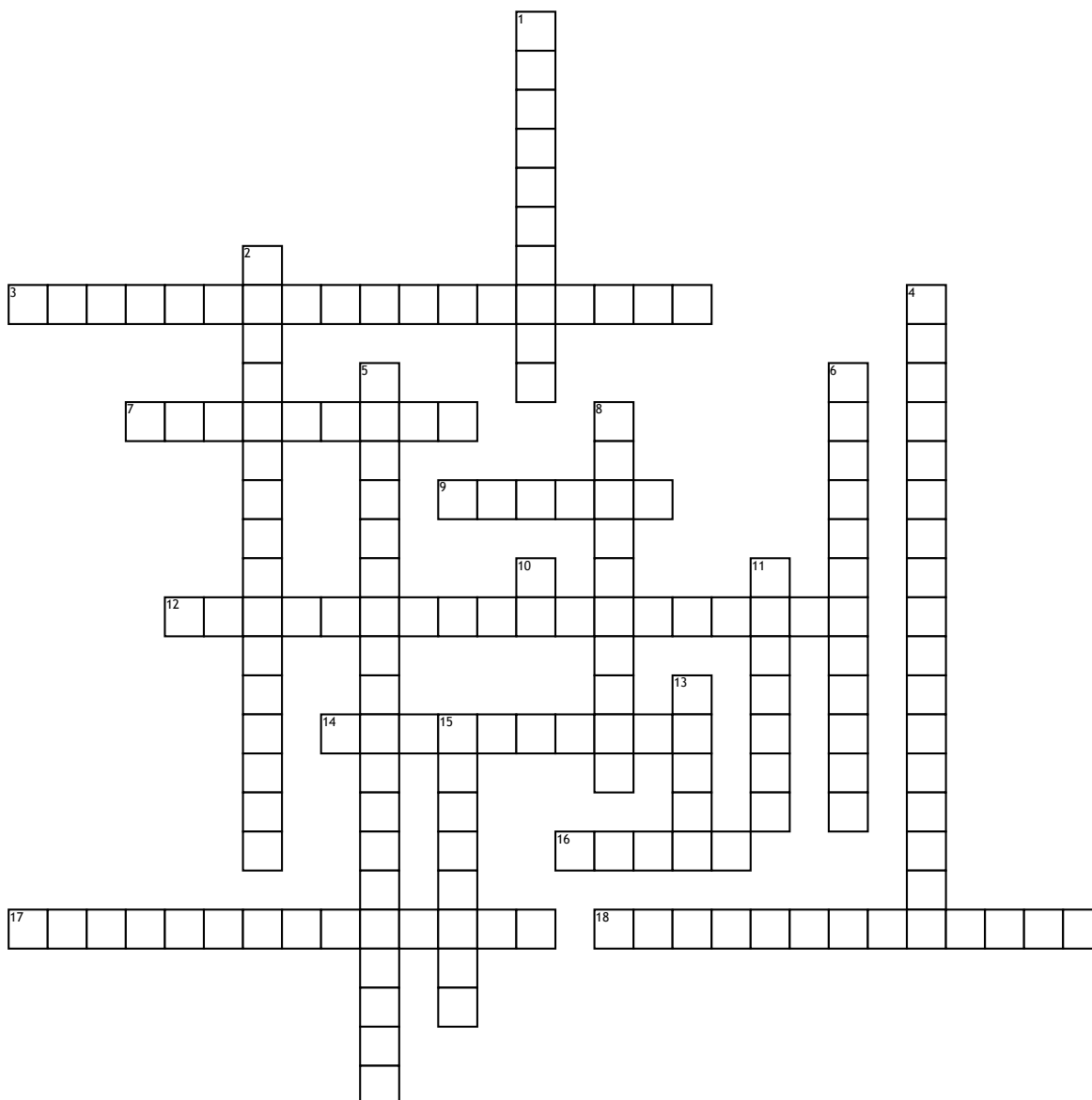


Covalent Bonding



Across

3. The name of N₂O

7. A condition that occurs when more than one valid Lewis structure can be written

9. Bond that forms when parallel orbitals overlap and share electrons

12. Occurs when more energy is released during product formation than is required to break bonds in the reactants

14. Single covalent bonds are also called

16. In naming binary acids this prefix is used

17. Electron-dot diagrams can be used to show valence electrons of atoms

18. The bond character that has a 0.4 - 1.7 electronegativity difference

Down

1. The strongest of the three types of covalent bonds

2. The bond character with 0 electronegativity difference

4. This molecular model uses letter symbols and bonds to show relative position of atoms

5. Occurs when a greater amount of energy is required to break the existing bonds in the reactants than is released when the new bonds form in the products

6. The chemical bond that results from sharing valence electrons

8. The longest of the three types of covalent bonds

10. In naming oxyacids the suffix "ate" changes to this suffix

11. An example of a covalent network solid

13. Valence Shell Electron Pair Repulsion

15. formed when two or more atoms bond covalently