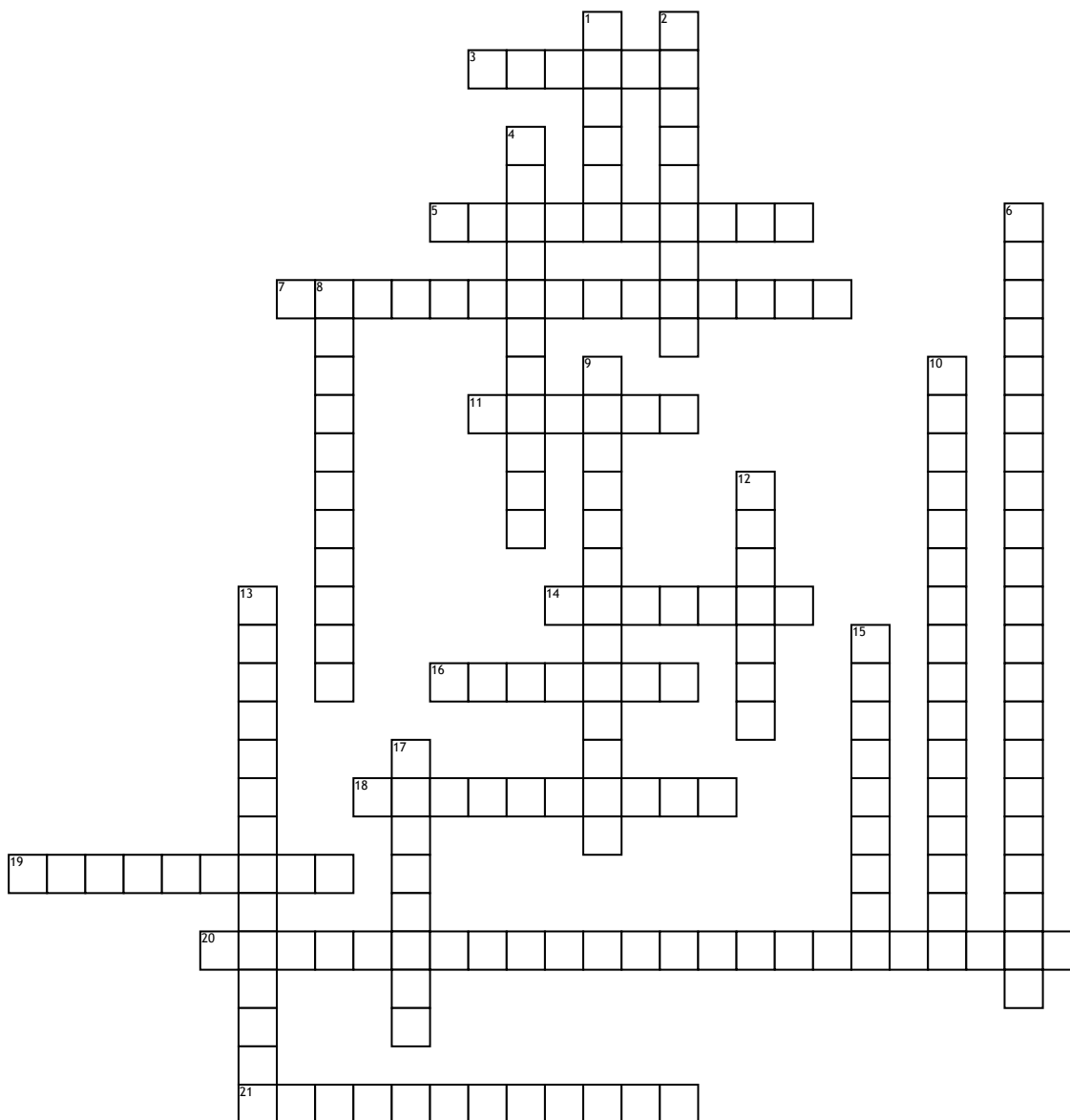


Name: _____

Date: _____

IPC - Chapter 10 Vocabulary



Across

3. A molecule that has both negatively and positively charged poles caused by the unequal distribution of electrons.
 5. A heterogeneous mixture consisting of small particles spread throughout a liquid or gaseous medium, from which they will eventually settle out.
 7. A method of determining the concentration of a solution by comparing the density of the solution to the density of water.
 11. A homogeneous mixture of metals.
 14. Dissolvable.
 16. The substance that does the dissolving in a solution.
 18. The maximum amount of a solute that can dissolve in a given amount of solvent under normal conditions.

19. Containing the maximum amount of a solute that can be dissolved in a given amount of solvent under normal conditions.

20. The effect whereby a solute lowers the freezing point of the solvent in which it is dissolved.

21. The process whereby a solvent breaks up an ionic solid.

Down

1. The substance that is dissolved in a solution.
 2. The greater the pressure on a liquid, the greater the amount of gas that will remain dissolved in that liquid at any given temperature.
 4. The property that allows two liquids to be soluble in each other.

6. The effect whereby a solute raises the boiling point of the solvent in which it is dissolved.

8. A solid formed during a reaction; insoluble in water

9. A molecule that has partially charged electrical poles.

10. A method of expressing the concentration of a solute as a percentage of the total mass of the solution.

12. A solution with a liquid solute and a solid solvent.

13. Having dissolved more than the normal amount of solute in a given amount of solvent.

15. Cannot be broken apart by water molecules (p 228)

17. a homogeneous mixture of two or more substances