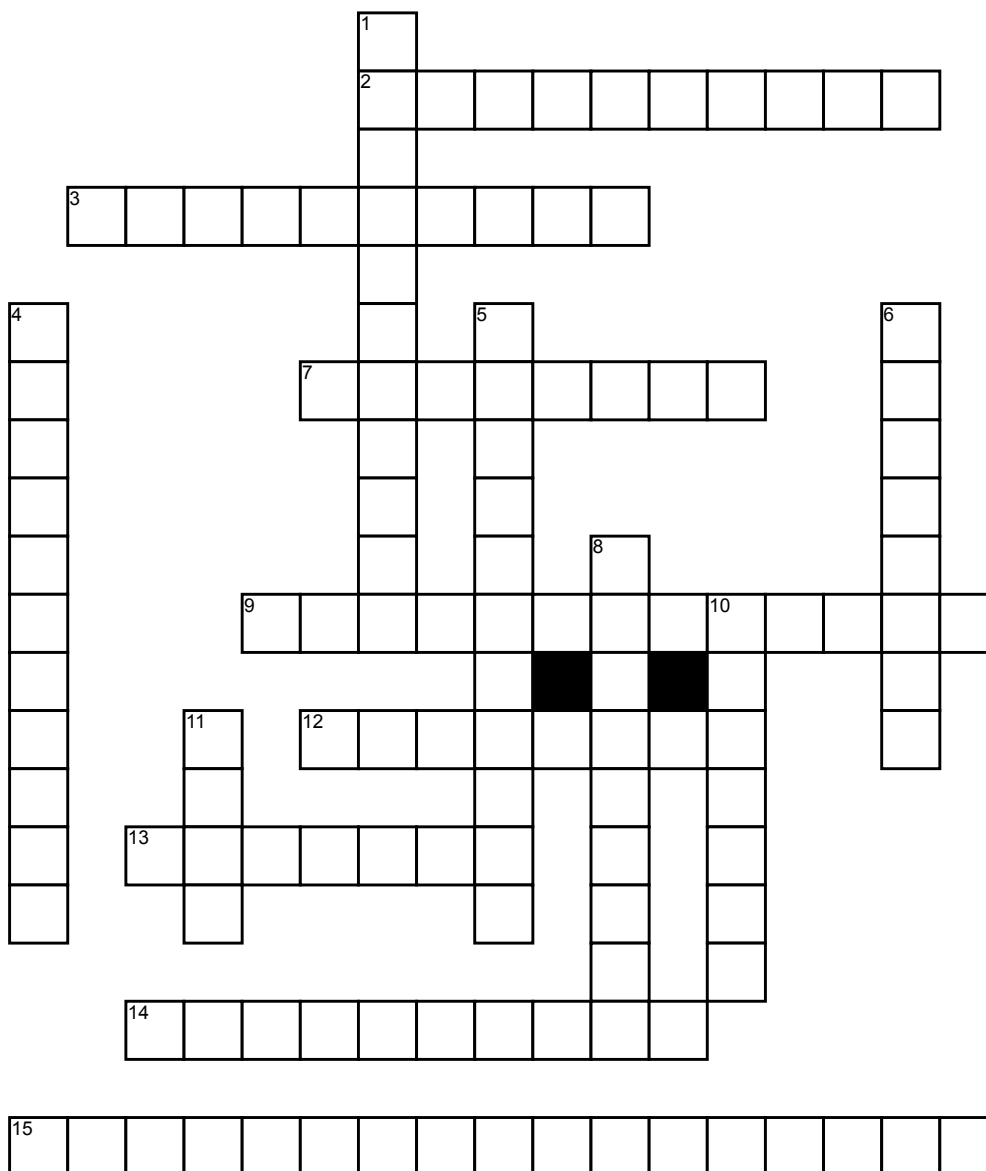


Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

# Introduction to the Periodic Table



## **Across**

- 2.** J.J. Thomson's classic experiment to determine the identity of a subatomic particle  
**3.** The first person to state that atoms were indivisible and indestructible  
**7.** Determined the mass and charge of an electron  
**9.** In Bohr's model of the atom, the electrons are in fixed energy levels that move in what type of path?

- 12.** Ernest Rutherford's experiment that determined a subatomic particle residing in the nucleus  
**13.** Credited for the discovery of the electron  
**14.** Identified that atoms were comprised of positive charges located in the nucleus of an atom  
**15.** The electron exhibits wave-particle behavior in this model of the atom.

## **Down**

- 1.** His mathematical descriptions solve for the characteristic shape and energy of electron density

- 4.** John Dalton's model of the atom is best described by what geometric shape  
**5.** J.J. Thomson's model of the atom  
**6.** Graduate student who found the neutron  
**8.** Transformed the way science was done by providing written observations, explanations and re-testable procedures.  
**10.** Millikan's historic experiment to determine the properties of a subatomic particle  
**11.** The planetary model of the atom was proposed by this scientist.