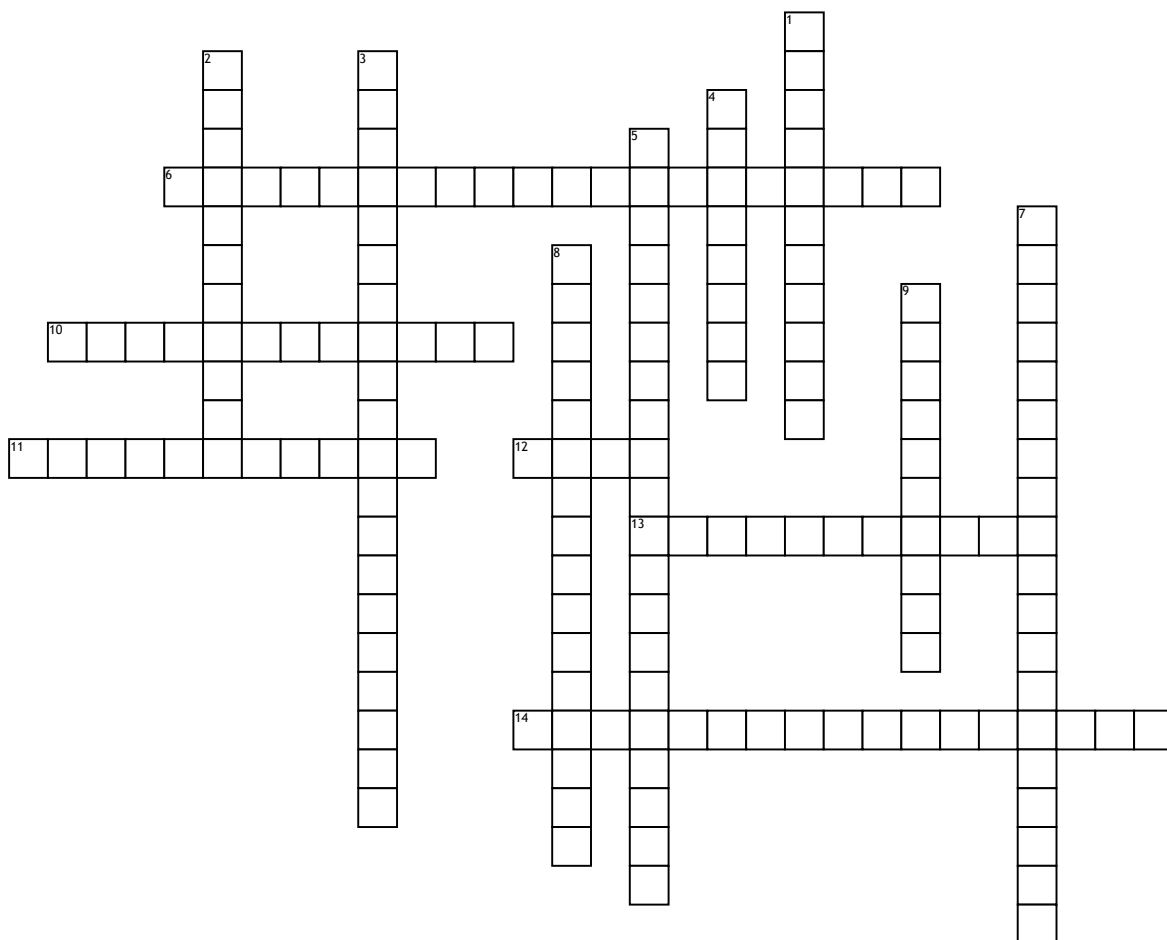


Name: _____

Date: _____

Kinetics



Across

6. States molecules must collide with enough force to break the bonds of the reactants and rearrange to form new substances

10. A chemical species that is a product in one step of a reaction mechanism and then a reactant in a subsequent step

11. Increases reaction events by allowing more reactant molecules to contact one another

12. The amount of product produced per time unit for a given reaction

13. Type of reaction for Rate = $k[A][B]$

14. A reaction that involves a series of reaction steps

Down

1. Increases the number of reaction events by giving reactant molecules more kinetic energy

2. Type of reaction where reactants have less energy than products

3. Products decided be reactions raised to the powers of the coefficients

4. Lowers the activation energy by creating a new pathway

5. States a reaction in equilibrium will shift to counteract the change imposed on it

7. The slowest step in a reaction mechanism

8. Minimum energy threshold for a reaction to occur

9. Type of reaction where products have lower energy than reactants