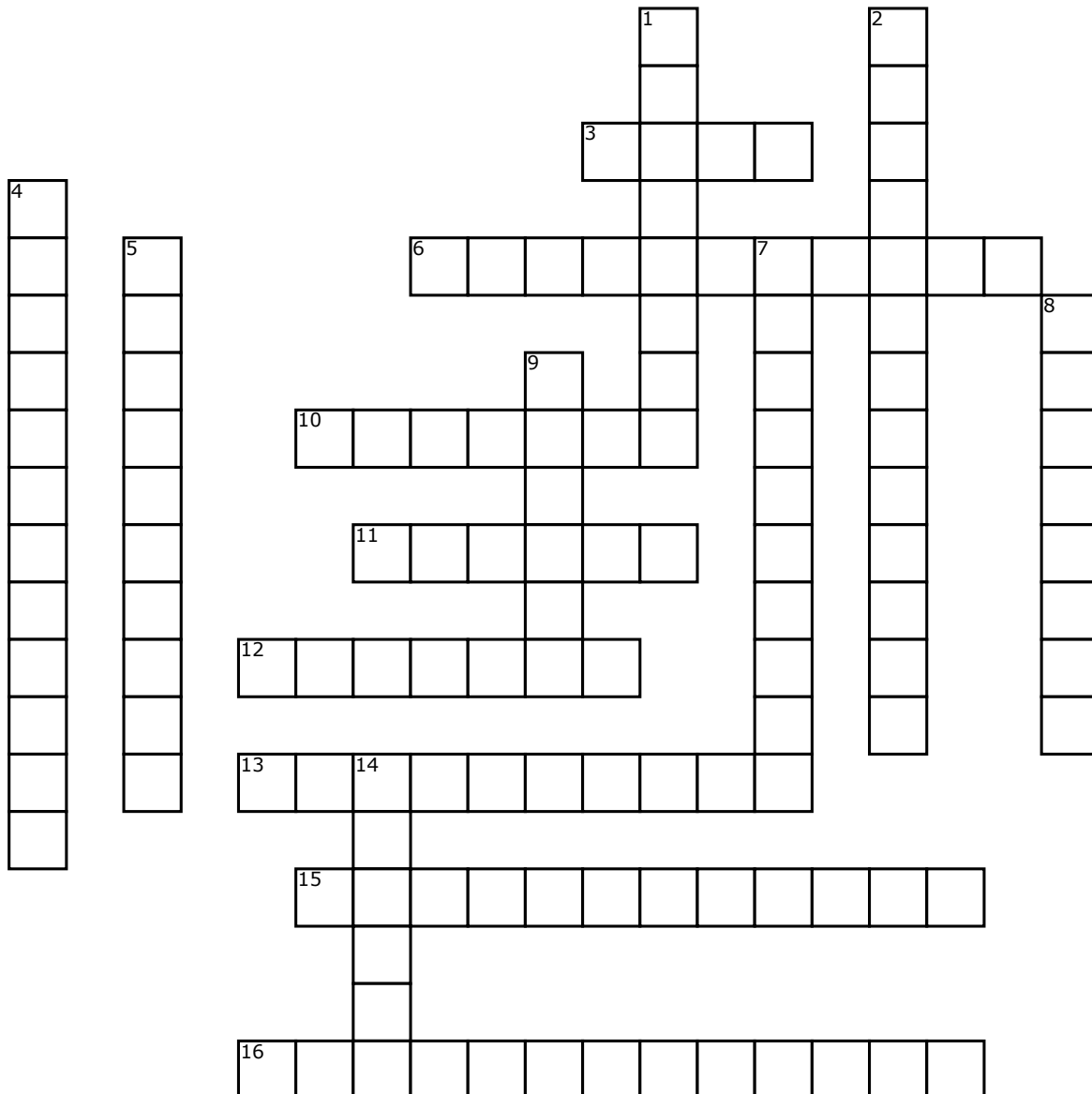


# Matter and Atoms



## **Across**

**3.** came up with a theory that said the electrons do not spiral into the nucleus and came up with some rules for what does happen.

**6.** J.J. Thompson's atom model's name

**10.** the neutral part of an atom

**11.** He performed experiments with various chemicals that showed that matter seemed to consist of elementary lumpy particles (atoms)

**12.** Can't be broken down any longer (an atom)

**13.** people who need to take notes

**15.** the whole number on the periodic table

**16.** the chart where all of the elements are

## **Down**

**1.** In 1897, the English physicist discovered the electron and proposed a model for the structure of the atom. he knew that electrons had a negative charge and thought that matter must have a positive charge

**2.** The people who are going to get an A for this

**4.** decimal on the periodic table

**5.** In 1911 he thought it would be interesting to bombard atoms with these alpha rays, figuring that this experiment could investigate the inside of the atom.

**7.** First person to think of the atom

**8.** the negative charge of an atom

**9.** the positive part of the atom

**14.** Everything/anything that takes up space