

Name: _____

Date: _____

Rates of Reaction

Across

2. Anything that increases the number of successful collisions between reactant particles will speed up a _____

5. The presence of a _____ increases the reaction rate

7. At a higher temperature, particles move _____

10. Increasing the _____ increases the speed of the molecules

12. A _____ reactant will react faster if it is broken into _____ pieces. This increases the _____ and allows more _____ between the reactants

13. Chemicals react together when their particles _____ with each other.

16. The speed of a chemical reaction may be defined as the change in concentration of a substance divided by the _____

17. the smaller a reactant the larger its _____

18. How quickly or slowly a reaction takes place

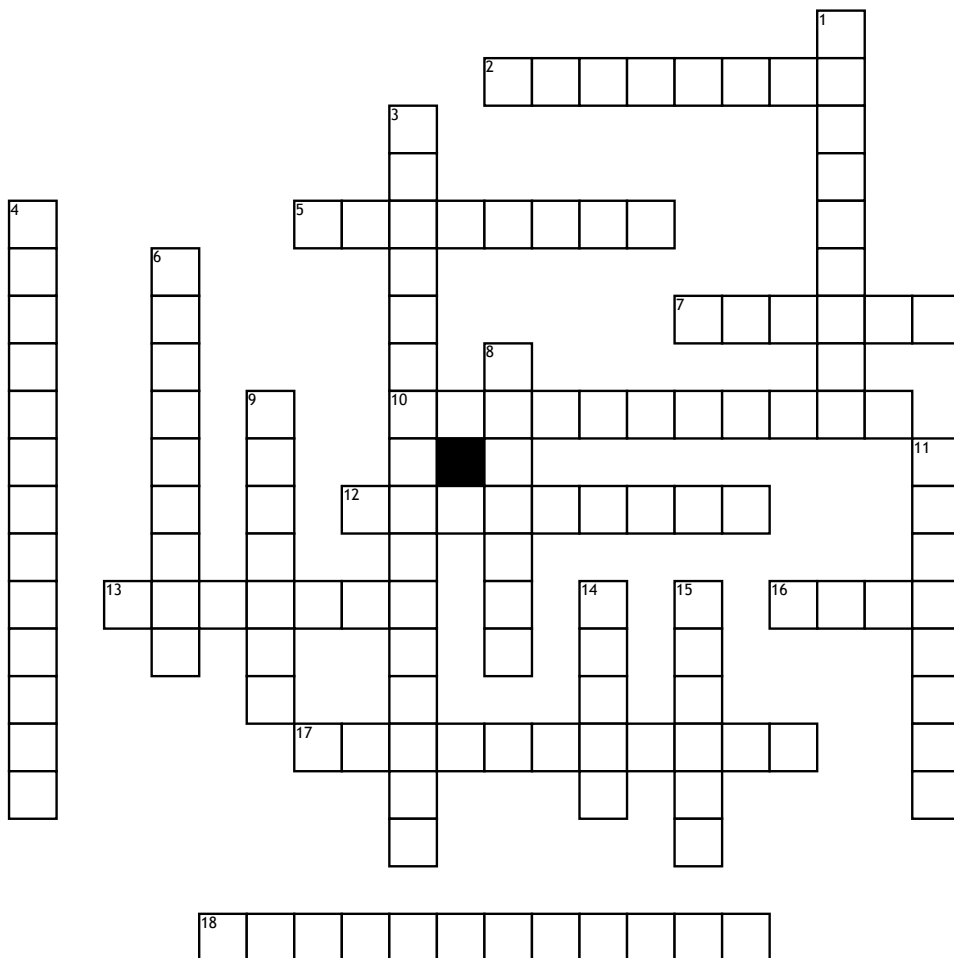
Down

1. A substance that blocks a catalyst

3. Adding a _____ can also speed up a reaction because it helps to lower the _____

4. At a higher _____, there are more particles in the same amount of space.

6. An increase in surface area means that there will be more _____ between the reactants.



8. A solid reactant will react faster if it is broken into _____ pieces.

9. As the pressure increases, the space in which the gas particles are moving becomes _____

11. More kinetic energy means more _____

14. The rate of reaction is used to describe the _____ of reaction.

15. particles must be moving very fast and have lots of kinetic _____ for collisions to occur.

