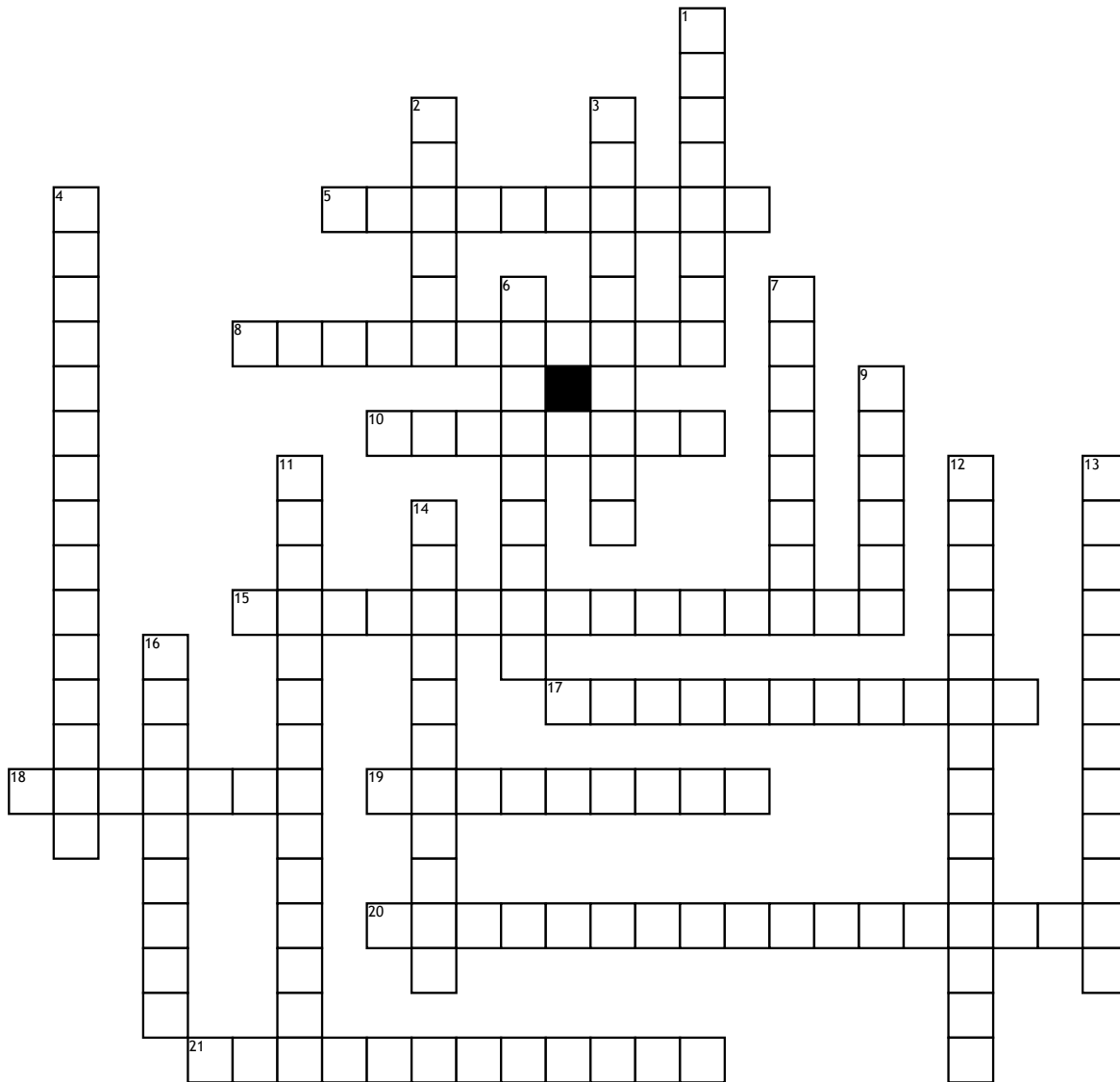


# Solution Vocabulary



## Across

5. temperature, pressure (for gases), surface area, stirring, and polarity (like dissolves like) are all factors that affect...  
 8. the difference in temperature between the boiling point of a solution and of the pure solvent  
 10. a homogeneous mixture  
 15. a solution in which only a fraction of the solute exists as ions  
 17. a solution that contains less solute than a saturated solution at a given temperature and pressure  
 18. the dissolving medium in a solution  
 19. a process that occurs when an ionic solute dissolves  
 20. a solution in which almost all of the solute exists as separate ions

21. the difference in temperature between the freezing point of a solution and of the pure solvent

## Down

1. the process of reducing the concentration of a solute in solution, usually simply by mixing with more solvent  
 2. a solution that contains a small amount of solute  
 3. the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution  
 4. a solution in which the solvent is water  
 6. a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure

7. the concentration of solute in a solution expressed as the number of moles of solute dissolved in 1 liter of solution  
 9. dissolved particles in a solution  
 11. a solution that contains more solute than a saturated solution at a given temperature and pressure  
 12. a compound that does not conduct an electric current in aqueous solution or in the molten state  
 13. a solution containing a large amount of solute  
 14. a compound that conducts an electric current in aqueous solution or in the molten state  
 16. stirring, increasing the surface area of the solute, changing temperature, and increasing pressure (for gases) are all ways to increase...