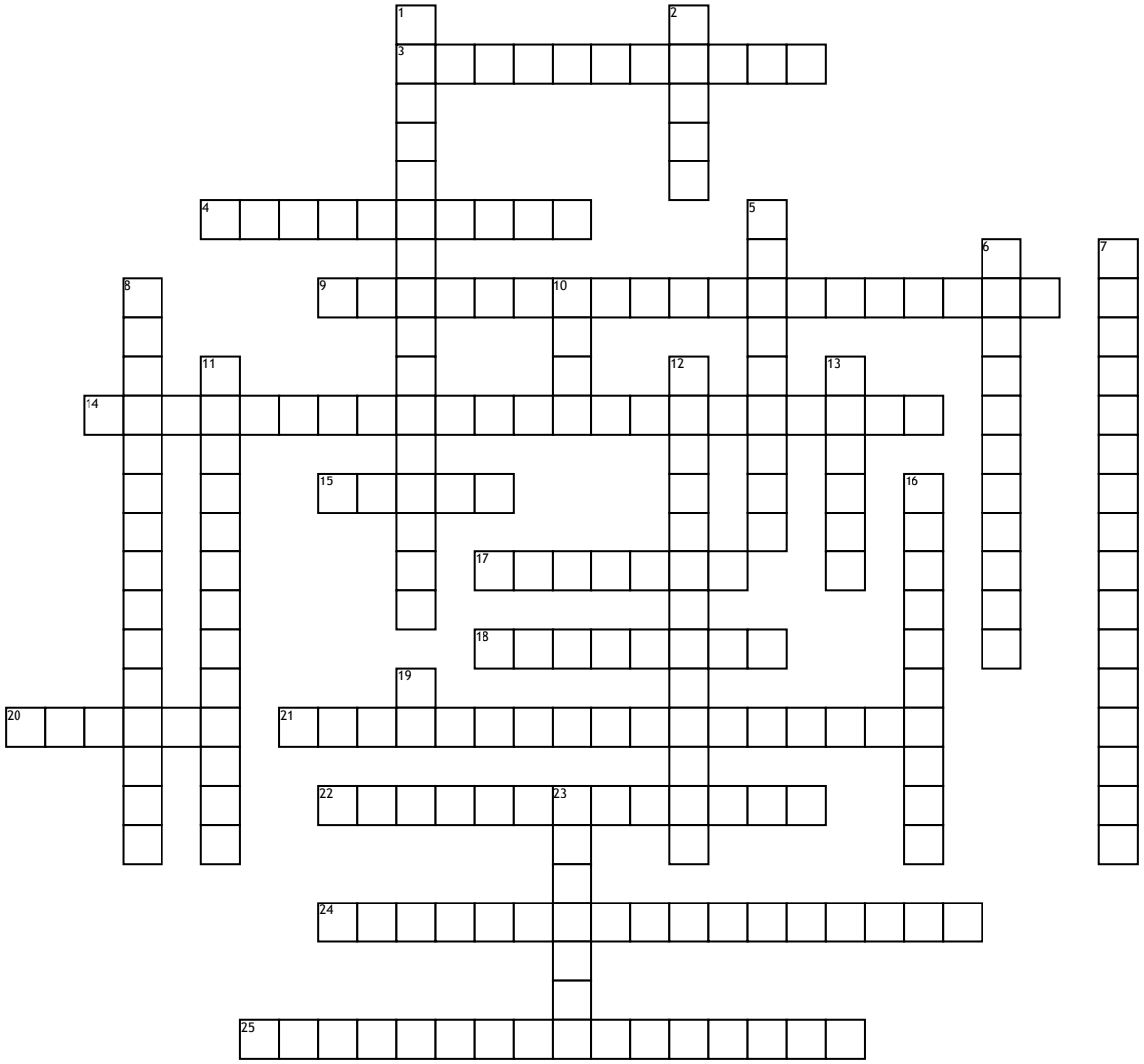
Name: _____ Date: _____

# Solutions, Acids, and Bases



**Across**

**3.** Ammonia, antacids, egg whites, and baking powder are examples of_____.

**4.** Maximum amount of a solute that can be dissolved in a given amount of solvent

**9.** Any solution that can dissolve more solute at a given temperature.

**14.** Solution that contains too much solute.

**15.** Produce hydrogen ions in water

**17.** Substance doing the dissolving

**18.** Physically mixed-sitting next to each other, but not joined chemically.

**20.** Substance that is dissolved.

**21.** Large amount of solute in a solution.

**22.** Increased concentrations of solute in a solution will lower the _____.

**24.** Solution that contains all the solute it can hold at a given temperature.

**25.** The time it takes a solute to dissolve in a solvent in a given solution

**Down**

**1.** Solutions are mad of H+ ion and an anion.

**2.** Produce hydroxide ions in water

**5.** Unable to be dissolved

**6.** Citric acid in citrus fruits

**7.** Water is called this because it dissolves more substances than others.

**8.** A curve on a graph that shows how much solute can be dissolved.

**10.** The time it takes an action to occur.

**11.** A mixture that does not have a uniform composition.

**12.** The amount of solute in a given amount of solvent.

**13.** Low amount of solute in solution.

**16.** A mixture that has a uniform composition.

**19.** Measurement of amount of hydrogen ions

**23.** The number 7 on the pH scale