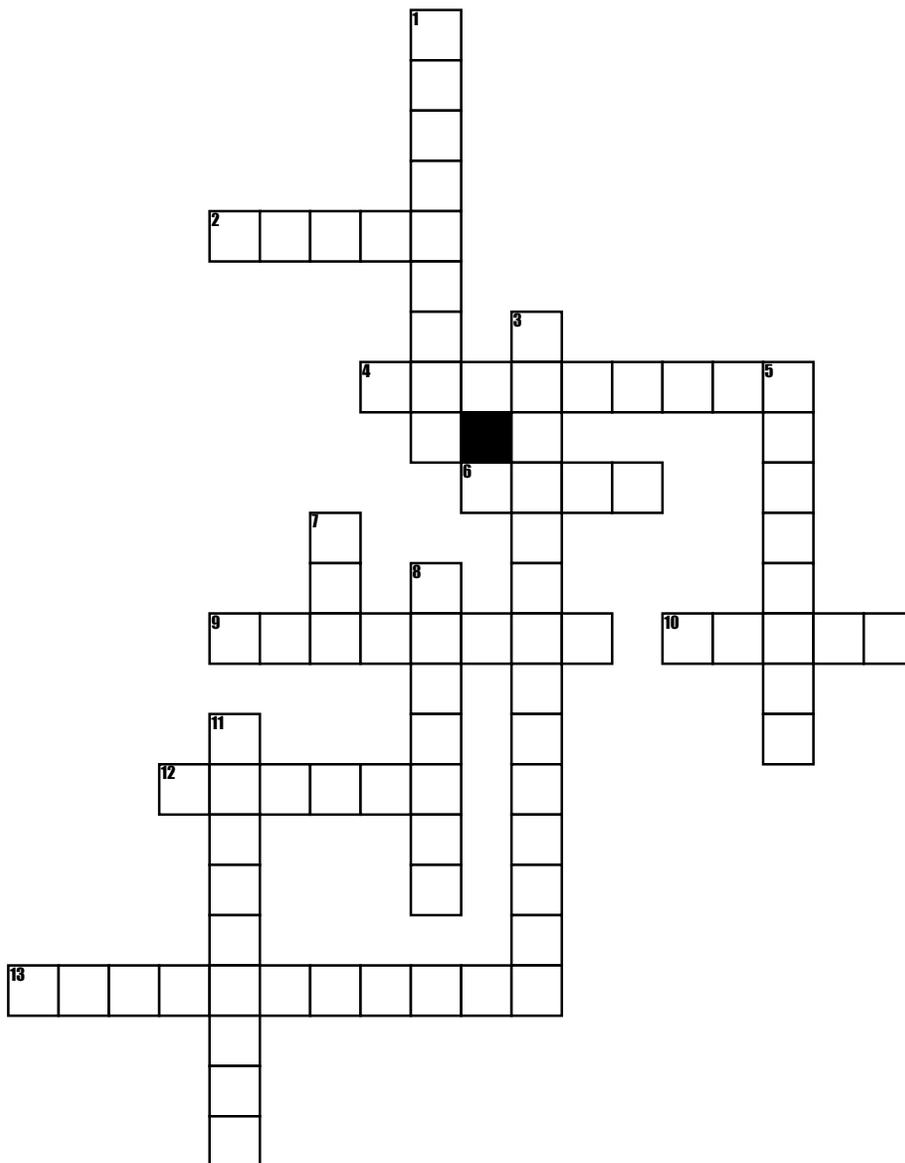


Solvation Process-chapter4-wolever,serissa



Across

- 2.** When a compound is held together by electrostatic forces and is composed of anions and cations, it is...
- 4.** When a solute no longer can dissolve any more solvent causing the excess to sink to the bottom
- 6.** A factor that can affect the solubility rate
- 9.** Nonpolar will dissolve...
- 10.** Polar will dissolve...
- 12.** The component that is being dissolved in the solvent

13. When a solute still has the ability to dissolve more solvent thus creating more solution it is...

Down

- 1.** When atoms share electrons through covalent bonds in a compound it is...
- 3.** When a substance is heated up after it has already hit its saturated point so that now the excess becomes completely dissolved but when it is later cooled off, it crystalizes it is...
- 5.** Process where molecules attract to one another to form a solution
- 7.** Aka molecule or atom. It obtains either a positive or negative charge and is created through ionization
- 8.** The substance that dissolves the solute
- 11.** Process of which the molecules of a solvent are attracted and associated with the molecules of a solute.