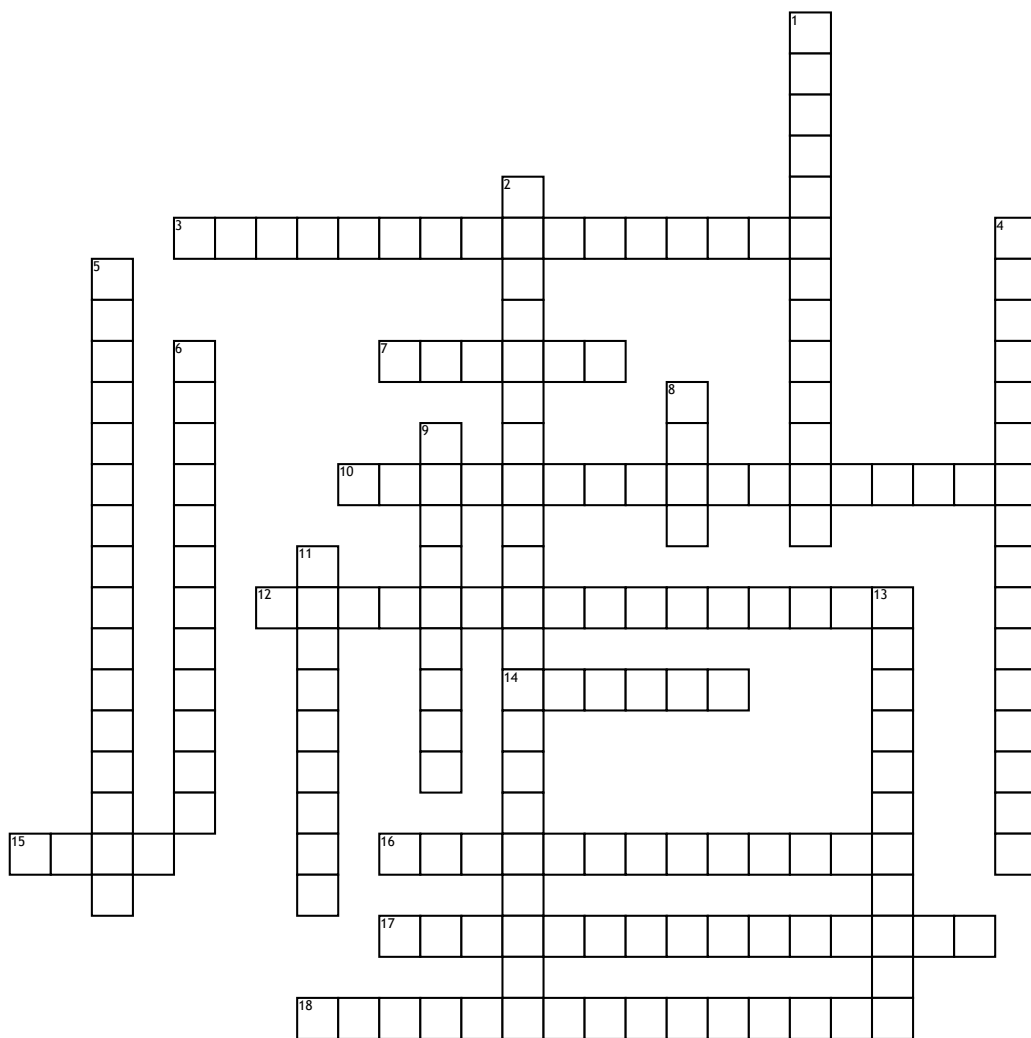


Name: _____

Date: _____

The Mole (Unit 5)



Across

3. (N) represents what?
 7. product of crude-oil refining
 10. The name of the value 6.02 times 10 to the 23
 12. A formula that shows the simplest whole number ratio of elements in a compound
 14. An alloy is a mixture of different what?
 15. A group of 6.02 entities (molecules, atoms, ions, electrons) is called a what?
 16. The study of the mass and amount relationships between reactants and products in a chemical reaction

17. the reactant that is completely consumed in a reaction. Determines how much product is made
 18. The ratio, expressed as a percentage, of the actual yield to the theoretical yield

Down

1. The reactant that is still present after the reaction is complete (left over)
 2. The percentage, by mass of each element in a compound
 4. The amount or mass of product predicted based on the stoichiometry of the chemical reaction

5. A formula that shows the element symbols and exact number of each type of atom in a molecular compound
 6. Measured in grams/mol (g/mol)
 8. (m) represents what?
 9. The ratio of the amounts of the entities in a chemical reaction
 11. The quantity of a substance measured in moles
 13. The amount or mass of product actually collected during an experiment or industrial process

Word Bank

Amount (n)
 molecular formula
 limiting reagent
 mass
 stoichiometry

butene
 excess reagent
 Mole
 Mole ratio
 percentage yield

Empirical formula
 Molar Mass (M)
 metals
 Avogadro's constant

percentage composition
 number of entities
 actual yield
 theoretical yield