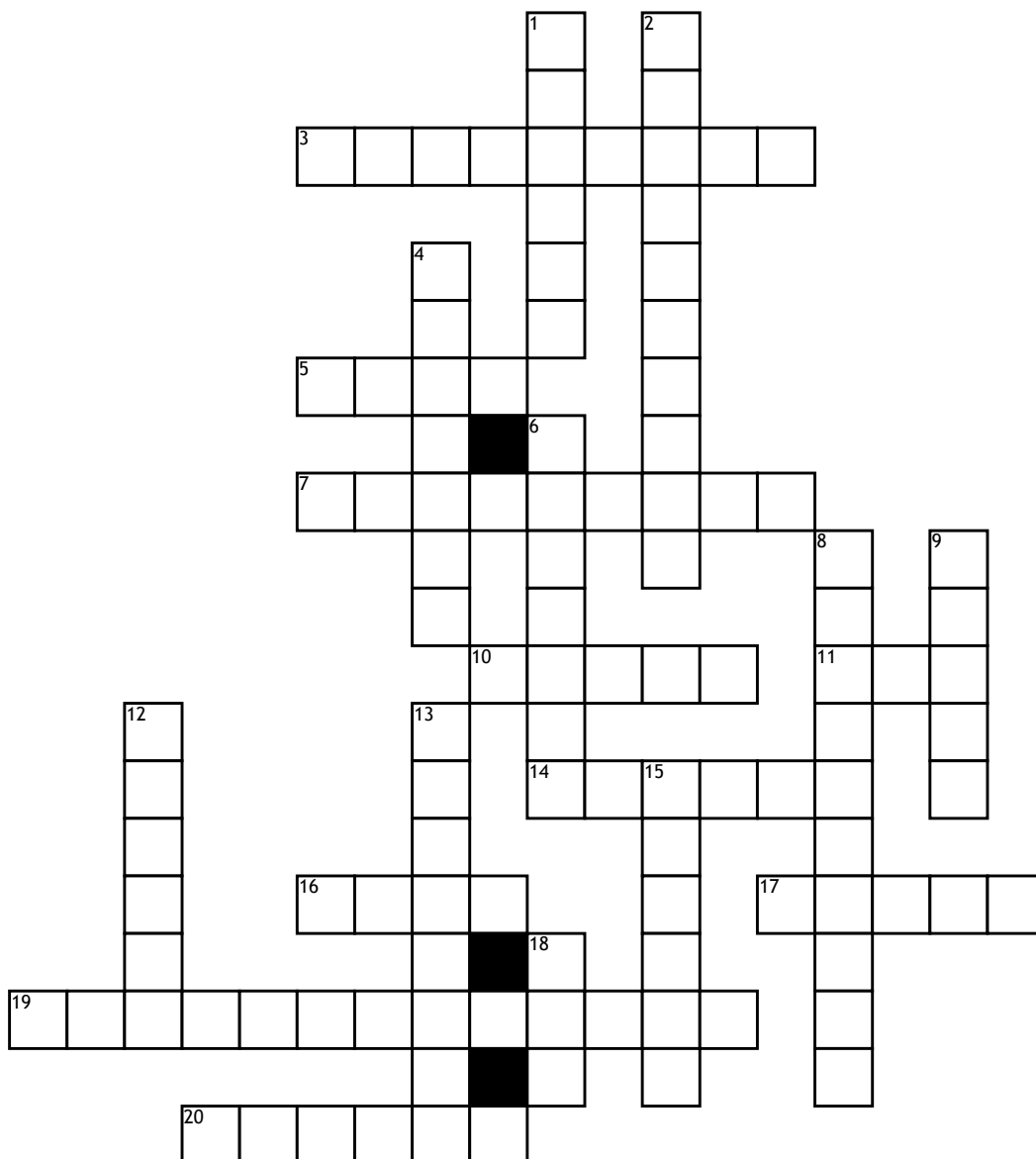


The Periodic Table



Across

3. _____ was the inventor of the table, he originally only had 60 elements.

5. The symbol Au is for _____ on the periodic table, Au stands for Aurum which is latin for _____.

7. Groups 13-16 are _____, they are dull, brittle, and don't conduct heat or electricity.

10. _____ gases have outer electron shells that are full, they don't often react with other elements.

11. Each element in period two has _____ valence electron shells.

14. The chemical _____ is one or two letters used to represent the elements name.

16. The atomic _____ is the average mass of an atom of an element.

17. One key physical difference between transition metals and poor metals is the _____ of the elements.

19. A model used for classifying elements and used to predict properties of elements.

20. Groups 1 and 2 are _____ metals and _____ earth metals, which are both very reactive.

Down

1. Periods describe the number of electron _____ that elements have.

2. Metal objects conduct an _____ current.

4. Electrons in the outer shell are _____ electrons.

6. The horizontal rows are called _____. There are 7 _____ on the table.

8. _____ have characteristics of nonmetals and metals.

9. A _____ is a vertical column on the periodic table. Elements in the same _____ share similar properties.

12. The atomic _____ is the number of protons in an atom of an element.

13. The first organization of the elements was based on _____ properties.

15. _____ make up almost 75% of the periodic table.

18. You can find elements with an atomic number over 92 in a _____.