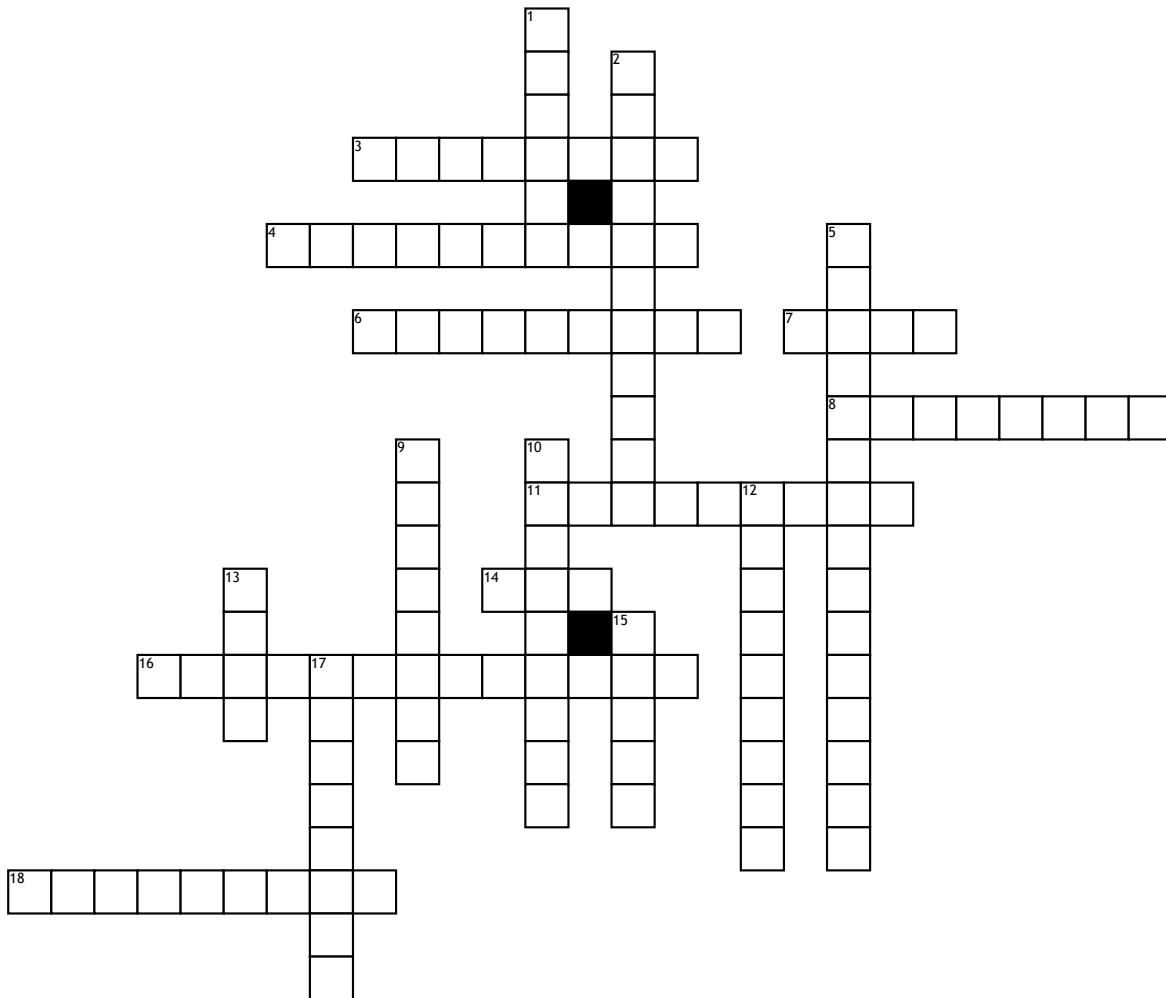


# The Scientific Mole



**Across**

3. To convert moles to mass, you have to \_\_\_\_\_ the measurements

4. The idea of the mole was first conceived as a \_\_\_\_\_ until proven correct as a law

6. One must use this other calculation of an element to determine the mole of a certain substance

7. The mass of substance containing the same number of fundamental units as there are atoms

8. The Italian Scientist who created the hypothesis that relates molar mass to physical mass

11. The mole is commonly used in this school subject

14. The unit of "mole" is commonly abbreviated as this other spelling

16. The mole is typically calculated in the conversion policy that was founded on the conservation of mass

18. The mole contributes to the \_\_\_\_\_ theory

**Down**

1. To convert mass to moles, you have to \_\_\_\_\_ the measurements

2. The term "mole" was first used to describe this property that depend on the ration between solute and solvent particles

5. Jean Baptiste Perrin was the first to coin this phrase which references the number Avogadro discovered

9. The mole is an SI unit that describes the \_\_\_\_\_ of a chemical entity

10. This unofficial holiday date celebrates the creation of Avogadro's number

12. The units used for a mole depend on the actual \_\_\_\_\_ used

13. A constituent particle refers to an molecule or an \_\_\_\_\_

15. The mole is typically expressed in this unit of measurement

17. Avagadro's Law is also known as Avogadro's \_\_\_\_\_