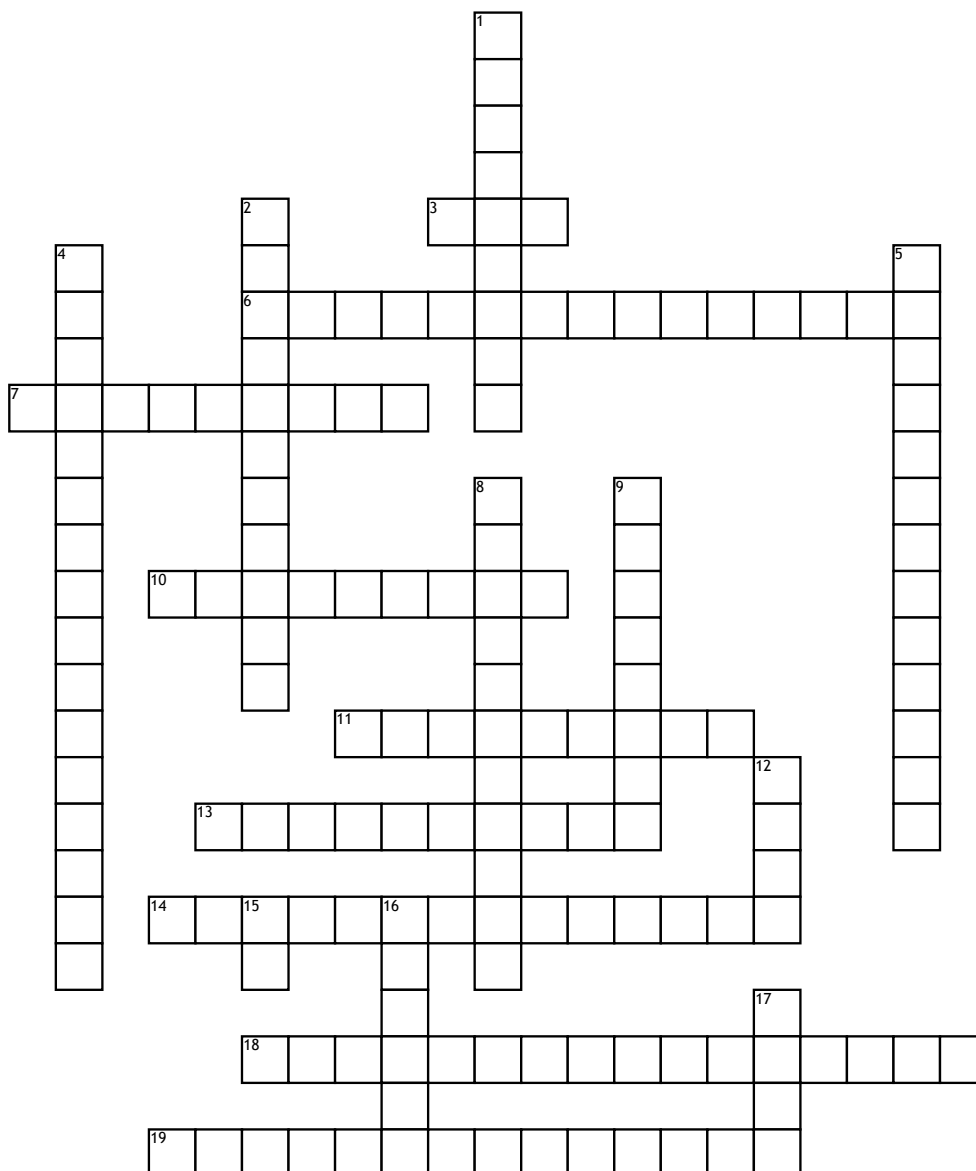


# Unit 9



## Across

3.  $-\log[\text{OH}^-] = \underline{\hspace{2cm}}$   
 6. shows the dependence of solubility on temperature  
 7. standard solution is used to determine the concentration of an unknown solution  
 10. holds the maximum amount of solute  
 11. acids form what kind of ions  
 13. max amount of solute that will dissolve in 100 g solvent at a given temp

14. contains more solute than the usual maximum amount and are unstable.

18.  $M_1V_1 = M_2V_2$

19. solution that does not conduct an electric current

## Down

1. bases form what kind of ions  
 2. amount of solute dissolved is less than the maximum that could be dissolved  
 4. minimum energy required for a reaction to occur

5. mass solute / mass solution x 100

8. solution that conducts an electric current

9. moles of solute / liter of solution

12. turn blue litmus red

15.  $-\log[\text{H}_3\text{O}^+] = \underline{\hspace{2cm}}$

16. substance being dissolved

17. turn red litmus blue