Name: _____ Date: _____

Vocabulary Quiz

1. average mass of one atom of an element (from particles in the nucleus)

2. the number of protons in the nucleus of an atom ; used to determine an elements position in the periodic table

3. a characteristic of a substance that describes how it combines with other substances to form new ones

4. a property of a subatomic particle ; positive (protons),negative (electrons), or neutral (nuetrons).

5. a negatively charged particle in the electron cloud surrounding the atomic nucleus

6. the negatively charged particle space containing electrons that surrounds the atomic nucleus

7. a pure substance that cannot be broken down chemically into simpler substances

8. vertical columns on the periodic table

9. a (neutral) particle with no electrical charge within the atomic nucleus

10. The positively charged center of an atom containing protons and neutrons

11. A conceptual model in which the elements are organized according to their properties; often displayed as a chart “periodic

12. The horizontal rows on the periodic table

13. A property of matter of matter that can be observed without changing the composition or identity of the matter “physical

14. A positively charged particle within the atomic nucleus; used to identify of an

15. Tendency of a substance to undergo chemical changes in a system

16. A particle smaller than an atom, such as a proton, neutron, or electron “subatomic

17. Electron located in the outer energy level (electron shell) “valence

18. The formation of a new substance with different properties; cannot be undone by physical means “chemical

A. reactivity

B. molecule

C. electron cloud

D. precipitate

E. coefficient

F. reaction”

G. nuetron

H. chemical property

I. property”

J. electron”

K. groups (families)

L. conservation”

M. “

N. change”

O. electron

P. table”

Q. atomic mass

R. formula”

19. A representation of a chemical reaction by symbols and numbers “chemical equation

20. A representation of a molecule or compound in which the elements are represented by their symbols and subscripts represents the number of atoms of each element “chemical

21. A change caused by the interaction of two or more substances resulting in the formation of new substances “chemical

22. The number placed in front of a chemical formula in a chemical equation; represents the number of molecules of that substance

23. A substance made of two or more elements

24. Matter is not created or destroyed; only rearranged “law of

25. Combined atoms of the same element

26. The formation of solids from a solution

S. compound

T. atomic number

U. nucleus

V. elements

W. electrical charge

X. element-proton

Y. particle”

Z. periods