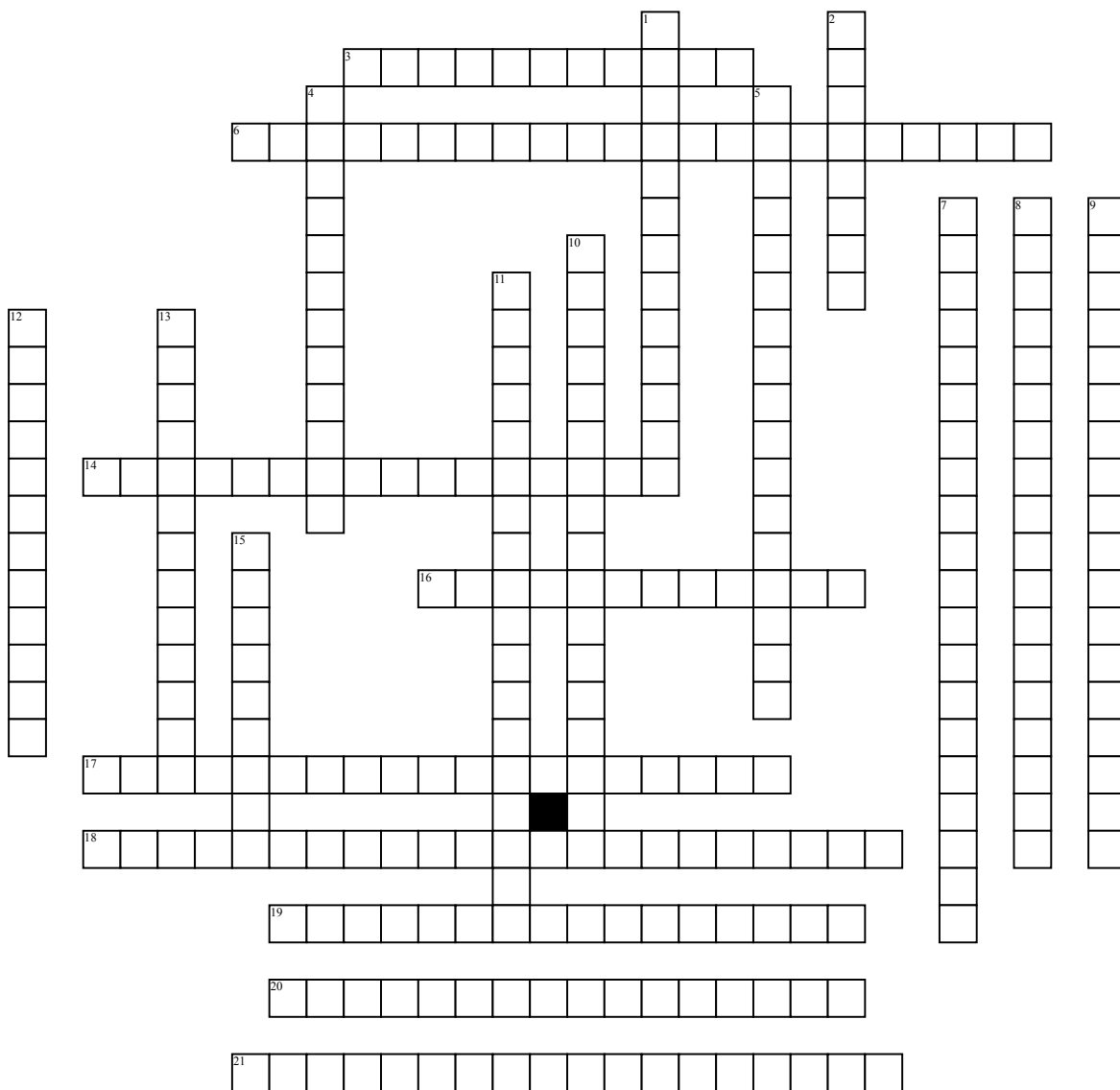


Name: _____

Date: _____

covalent bonding words



Across

3. used in chemistry to predict the geometry of individual molecules from the number of electron pairs surrounding their central atoms.

6. one measure of the strength of a chemical bond.

14. are molecules composed of only two atoms, of the same or different chemical elements

16. type of chemical bond where two atoms are connected to each other by the sharing of two or more electrons.

17. are two forms of a molecule where the chemical connectivity is the same but the electrons are distributed differently around the structure.

18. between two atoms that is produced when one atom shares a pair of electrons with another atom lacking such a pair.

19. a central atom is located at the center with four substituents that are located at the corners of a tetrahedron.

20. formula giving the number of atoms of each of the elements present in one molecule of a specific compound.

21. where two pairs of electrons are shared between the atoms rather than just one pair.

Down

1. Also called dipole

2. a group of atoms bonded together, representing the smallest fundamental unit of a chemical compound that can take part in a chemical reaction.

4. a pair of valency electrons of opposite spin that are not shared between the atoms in a molecule and are responsible for the formation of coordinate bonds.

5. compound where the atoms share electrons through covalent bonds.

7. are a type of chemical bond where two atoms share a pair of electrons with each other.

8. is a chemical bond between two atoms involving six bonding electrons instead of the usual two in a covalent single bond.

9. result when two dipolar molecules interact with each other through space.

10. formula that shows the arrangement of atoms in the molecule of a compound.

11. is when only one pair of electrons is shared between atoms.

12. a weak bond between two molecules resulting from an electrostatic attraction between a proton in one molecule and an electronegative atom in the other.

13. a charged chemical species (ion) composed of two or more atoms covalently bonded or of a metal complex that can be considered to be acting as a single unit.

15. is a covalent bond between two atoms where the electrons forming the bond are unequally distributed