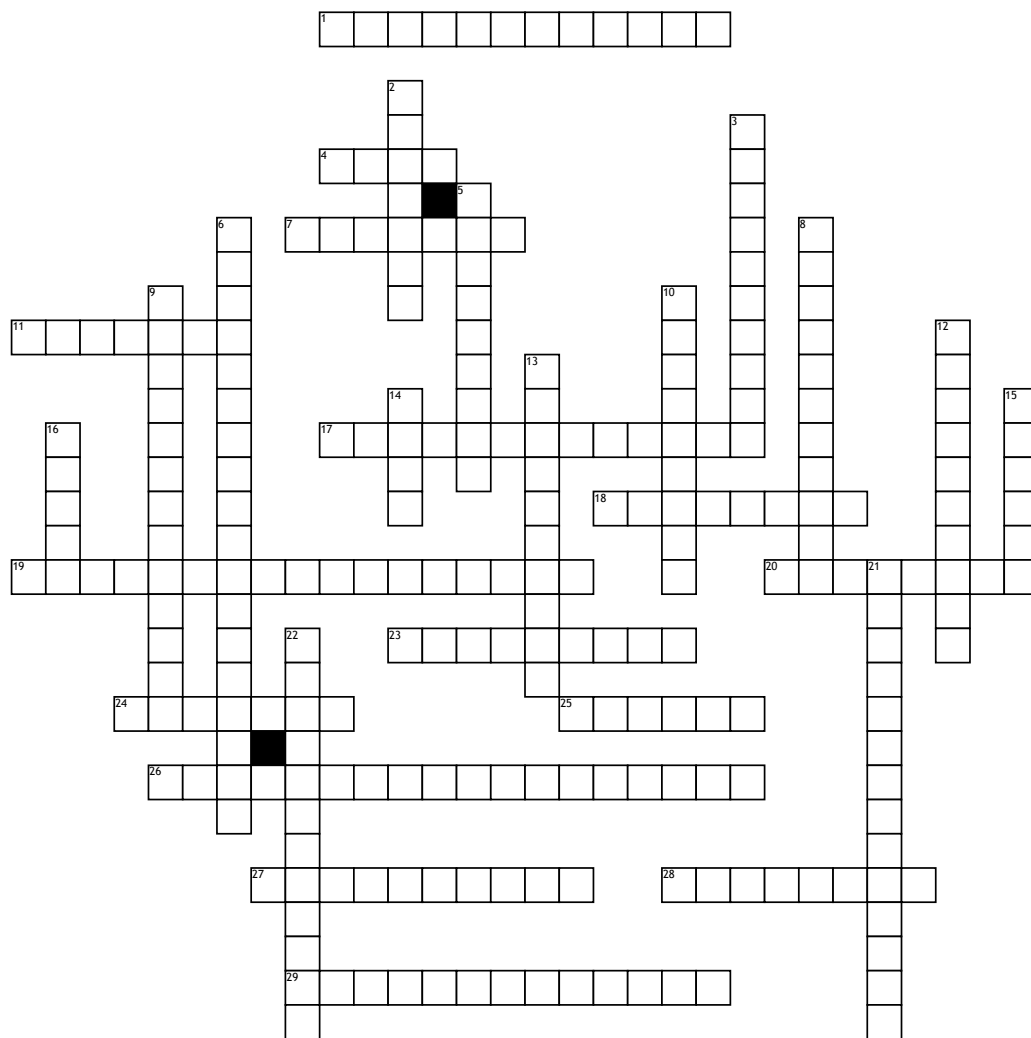


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# periodic table project



## Across

1. is a measure of the size of its atoms,
4. the basic unit of a chemical element.
7. any of the elements fluorine, chlorine, bromine, iodine, and astatine, occupying group VIIA (17) of the periodic table.
11. the central and most important part of an object, movement, or group, forming the basis for its activity and growth.
17. s is the small, dense region consisting of protons and neutrons at the center of an atom,
18. a single substance
19. is a measure of the tendency of an atom to attract a bonding pair of electrons
20. a stable subatomic particle with a charge of negative electricity, found in all atoms and acting as the primary carrier of electricity in solids.
23. a thing that is composed of two or more separate elements; a mixture
24. a subatomic particle of about the same mass as a proton but without an electric charge,
25. a stable subatomic particle occurring in all atomic nuclei, with a positive electric charge equal in magnitude to that of an electron, but of opposite sign.

26. . a property or characteristic of a substance that is observed during a reaction in which the chemical composition or identity of the substance is changed:

27. New Zealand-born British physicist who classified radiation into alpha, beta, and gamma types and discovered the atomic nucleus.

28. any of the gaseous elements helium, neon, argon, krypton, xenon, and radon, occupying Group 0 (18) of the periodic table.

29. . It is used to describe where electrons are when they go around the nucleus of an atom.

## Down

2. each of two or more forms of the same element that contain equal numbers of protons but different numbers of neutrons in their nuclei

3. the mass of an atom of a chemical element expressed in atomic mass units.

5. an element whose properties are intermediate between those of metals and solid nonmetals.

6. is measurable, whose value describes a state of a physical system

8. any of the elements lithium, sodium, potassium, rubidium, cesium, and francium, occupying Group IA (1) of the periodic table.

9. a table of the chemical elements arranged in order of atomic number, usually in rows, so that elements with similar atomic structure

10. He formulated the Periodic Law, created a farsighted version of the periodic table of elements,

12. the electrode of an electrochemical cell at which reduction occurs

13. is the process by which an atom or a molecule acquires a negative or positive charge by gaining or losing electrons to form ions, often in conjunction with other chemical changes.

14. He formulated the Periodic Law, created a farsighted version of the periodic table of elements,

15. the theory that matter consists of indivisible particles called atom

16. a solid material that is typically hard, shiny, malleable, fusible, and ductile, with good electrical and thermal conductivity

21. is a column of elements in the periodic table of the chemical elements

22. the number of protons in the nucleus of an atom, which determines the chemical properties of an element and its place in the periodic table.

## Word Bank

electron cloud	rutherford	isotope	proton	electron
alkali metal	atomic nucleus	atomic radius	bohr	metalloid
chemical properties	atom	atomic number	physical properties	ionization
metal	cathode ray	periodic table	dalton	nucleus
electronegativity	atomic mass	elements	compounds	noble gas
mendeleev	neutron	halogen	chemical family	