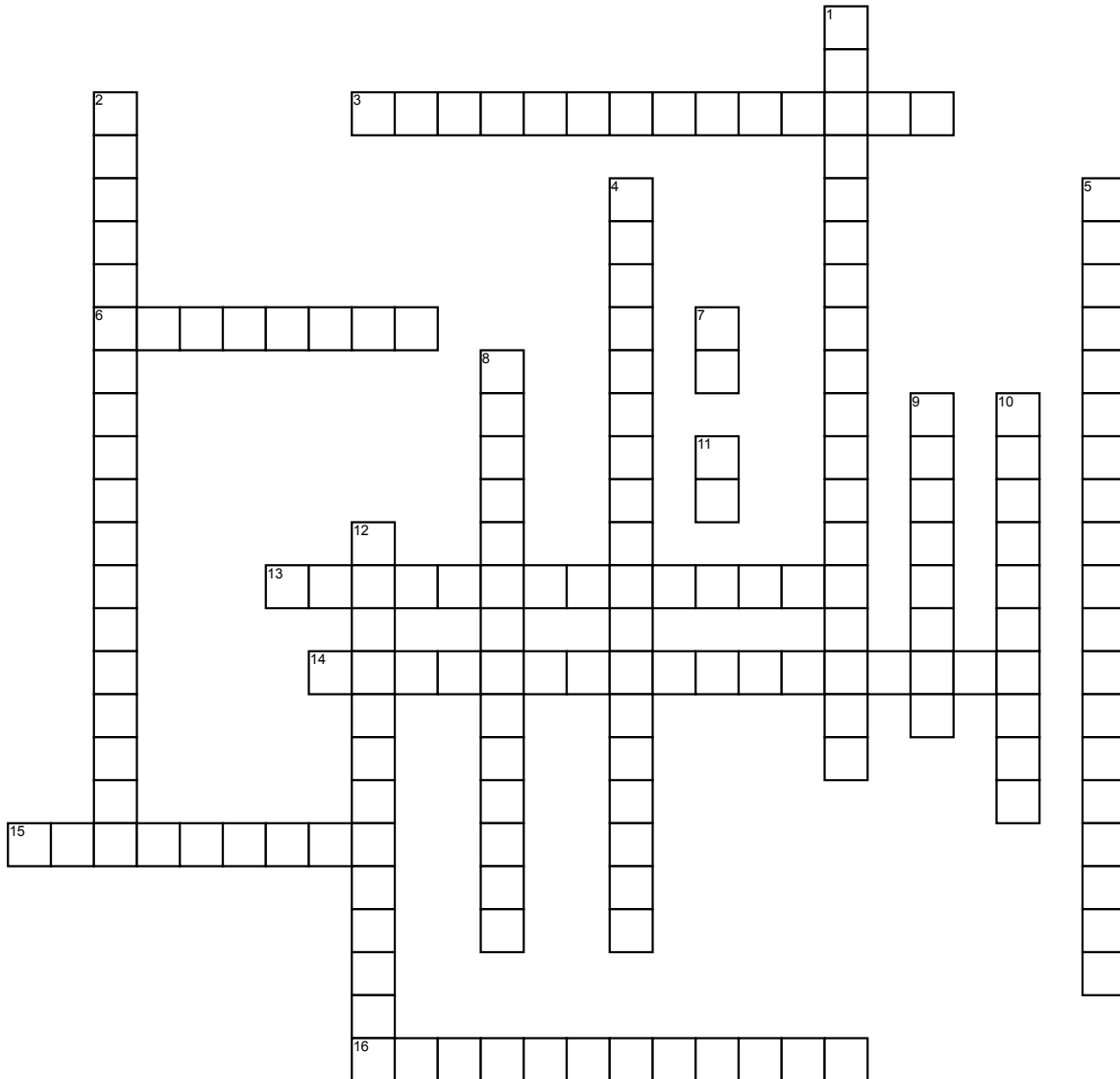


Name: \_\_\_\_\_

# unit 5 exam



## Across

3. fireworks
6. how many moles are in 40 grams of h<sub>2</sub>O
13. burning paper
14.  $2\text{NaBr} + \text{Ca}(\text{OH})_2 = \text{CaBr}_2 + \text{NaOH}$
15. how many grams are in 0.0106 moles of NO<sub>2</sub>
16. sodium + chlorine = sodium chloride

## Down

1.  $2\text{AgNO}_3 + \text{Cu} = \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$
2. magnesium carbonate = magnesium oxide + carbon Dioxide
4. burning sugar
5. cooking an egg
7. using the information from 8, my percent yield is 75%, how many grams of water did I make

8. a bike rusting
9. iron + oxygen yields iron oxide
10.  $\text{P}_4 + 3\text{O}_2 = 2\text{P}_2\text{O}_3$
11. if I start with 5 grams of C<sub>3</sub>H<sub>8</sub>, what is my theoretic yield of water
12.  $2\text{NO}_2 = 2\text{O}_2 + \text{N}_2$